

Name: _____



Balancing Chemical Equations

Directions: Balance each of the following chemical equations.

- 1) $\underline{\hspace{1cm}} \text{C}_{30}\text{H}_{22} + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 2) $\underline{\hspace{1cm}} \text{Al}(\text{OH})_3 + \underline{\hspace{1cm}} \text{HBr} \rightarrow \underline{\hspace{1cm}} \text{AlBr}_3 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 3) $\underline{\hspace{1cm}} \text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 4) $\underline{\hspace{1cm}} \text{C}_3\text{H}_8 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 5) $\underline{\hspace{1cm}} \text{Li} + \underline{\hspace{1cm}} \text{AlCl}_3 \rightarrow \underline{\hspace{1cm}} \text{LiCl} + \underline{\hspace{1cm}} \text{Al}$
- 6) $\underline{\hspace{1cm}} \text{C}_3\text{H}_8 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 7) $\underline{\hspace{1cm}} \text{NH}_4\text{OH} + \underline{\hspace{1cm}} \text{H}_3\text{PO}_4 \rightarrow \underline{\hspace{1cm}} (\text{NH}_4)_3\text{PO}_4 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 8) $\underline{\hspace{1cm}} \text{Rb} + \underline{\hspace{1cm}} \text{P} \rightarrow \underline{\hspace{1cm}} \text{Rb}_3\text{P}$
- 9) $\underline{\hspace{1cm}} \text{CH}_4 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 10) $\underline{\hspace{1cm}} \text{Al}(\text{OH})_3 + \underline{\hspace{1cm}} \text{H}_2\text{SO}_4 \rightarrow \underline{\hspace{1cm}} \text{Al}_2(\text{SO}_4)_3 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 11) $\underline{\hspace{1cm}} \text{Na} + \underline{\hspace{1cm}} \text{Cl}_2 \rightarrow \underline{\hspace{1cm}} \text{NaCl}$
- 12) $\underline{\hspace{1cm}} \text{Rb} + \underline{\hspace{1cm}} \text{S}_8 \rightarrow \underline{\hspace{1cm}} \text{Rb}_2\text{S}$
- 13) $\underline{\hspace{1cm}} \text{H}_3\text{PO}_4 + \underline{\hspace{1cm}} \text{Ca}(\text{OH})_2 \rightarrow \underline{\hspace{1cm}} \text{Ca}_3(\text{PO}_4)_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 14) $\underline{\hspace{1cm}} \text{NH}_3 + \underline{\hspace{1cm}} \text{HCl} \rightarrow \underline{\hspace{1cm}} \text{NH}_4\text{Cl}$
- 15) $\underline{\hspace{1cm}} \text{Li} + \underline{\hspace{1cm}} \text{H}_2\text{O} \rightarrow \underline{\hspace{1cm}} \text{LiOH} + \underline{\hspace{1cm}} \text{H}_2$
- 16) $\underline{\hspace{1cm}} \text{Ca}_3(\text{PO}_4)_2 + \underline{\hspace{1cm}} \text{SiO}_2 + \underline{\hspace{1cm}} \text{C} \rightarrow \underline{\hspace{1cm}} \text{CaSiO}_3 + \underline{\hspace{1cm}} \text{CO} + \underline{\hspace{1cm}} \text{P}$
- 17) $\underline{\hspace{1cm}} \text{NH}_3 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{N}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 18) $\underline{\hspace{1cm}} \text{FeS}_2 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{Fe}_2\text{O}_3 + \underline{\hspace{1cm}} \text{SO}_2$

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RJ Alexander



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