This activity will help you become skilled at balancing chemical equations.

- 1. Cut out the squares that the bottom of the paper.
- Decide whether or not each equation is balanced. If it is not, decide how to balance it, pasting the small squares as needed into the spaces.
- 3. Some spaces will not require a number.
- S₈ + ___ O₂ → ___ SO₃
- 2. ____ Na + ____ H₂O → ____ NaOH + ____ H₂
- 3. ___C₁₀H₁₆ + ___Cl₂ → ___C + ___HCl
- 4. ___ Si₂H₃ + ___ O₂ → ___ SiO₂ + ___ H₂O
- FeS₂ + ___ O₂ → ___ Fe₂O₃ + ___ SO₂
- 6. ___ Fe₂O₂ + ___ H₂ \rightarrow ___ Fe + ___ H₂O
- 7. ___ H₂O₂ → ___ H₂O + ___ O₂
- 8. ___ SiO₂ + ___ HF → ___ SiF₄ + ___ H₂O
- 9. ___ K + ___ Br₂ → ___ KBr
- 10. ___S₈ + ___ F₂ → ___ SF₆
- 4 11 2 2 2 8 2 16 2 6
 - 8 8 4 11 2 2 2 2 2 8 10
- 12 2 2 4 2 24 8

Balancing Act Practice Worksheet Answer Key

Jayne Bartlett

Balancing Act Practice Worksheet Answer Key:

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