Name: \_\_\_\_\_

Date: \_\_\_\_\_

## **Balancing Chemical Equations**

Balance the following chemical equations.

1. 
$$\underline{2}$$
 Fe +  $\underline{3}$  H<sub>2</sub>SO<sub>4</sub>  $\longrightarrow$   $\underline{1}$  Fe<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> +  $\underline{3}$  H<sub>2</sub>

2. 
$$\underline{1}$$
 CH<sub>4</sub> +  $\underline{2}$  O<sub>2</sub>  $\longrightarrow$   $\underline{1}$  CO<sub>2</sub> +  $\underline{2}$  H<sub>2</sub>O

3. 
$$1$$
 SiCl<sub>4</sub>( $\ell$ ) +  $2$  H<sub>2</sub>O( $\ell$ )  $\longrightarrow$   $1$  SiO<sub>2</sub>( $s$ ) +  $4$  HCl(aq)

4. 
$$\underline{2}$$
 AgI +  $\underline{1}$  Na<sub>2</sub>S  $\longrightarrow$   $\underline{1}$  Ag<sub>2</sub>S +  $\underline{2}$  NaI

5. 
$$\underline{4}$$
 NH<sub>3</sub> +  $\underline{5}$  O<sub>2</sub>  $\longrightarrow$   $\underline{4}$  NO +  $\underline{6}$  H<sub>2</sub>O

6. 1 FeO<sub>3</sub>(s) + 3 CO(g) 
$$\longrightarrow$$
 1 Fe( $\ell$ ) + 3 CO<sub>2</sub>(g)

7. 
$$\underline{1}$$
 SiO<sub>2</sub> +  $\underline{4}$  HF  $\longrightarrow$   $\underline{1}$  SiF<sub>4</sub> +  $\underline{2}$  H<sub>2</sub>O

9. 
$$4 (NH_4)_3PO_4 + 3 Pb(NO_3)_4 \longrightarrow 1 Pb_3(PO_4)_4 + 12 NH_4NO_3$$

# **Balancing Chemical Skeleton Equations Answer Key**

**David Thurlo** 

# **Balancing Chemical Skeleton Equations Answer Key:**

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study essential for understanding and developing new pharmaceutical substances The evolution of Pharmaceutical Organic Chemistry stems from its application in drug development integrating knowledge from organic chemistry into practical uses for pharmaceuticals Organic chemistry provides the foundation for biochemistry which explores health and disease and is critical for the practice of nutritional medical and related life sciences It also underpins advancements in medicinal chemistry bioinformatics biotechnology gene therapy pharmacology pathology chemical engineering dental science and more The **Language of Chemistry** Ruth A. Walker, Helen Johnston, 1967 Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 Addison-Wesley Chemistry ,1990 General Chemistry Ralph H. Petrucci, William S. Harwood, 1993 General Chemistry Principles and Modern Applications is recognized for its superior problems lucid writing and precision of argument This updated and expanded edition retains the popular and innovative features of previous editions including Feature Problems follow up Integrative and Practice Exercises to accompany every in chapter Example and Focus On application boxes as well as new Keep in Mind marginal notes Topics covered include atoms and the atomic theory chemical compounds and reactions gases Thermochemistry electrons in atoms chemical bonding liquids solids and intermolecular forces chemical kinetics principles of chemical equilibrium acids and bases electrochemistry representative and transitional elements and nuclear and organic chemistry For individuals interested in a broad overview of chemical **Chemistry** James C. Hill, 2003 This book assists students through the text material with principles and applications chapter overviews learning objectives review of key terms cumulative chapter review guizzes and self tests Included are answers to all Student Guide exercises Chapter summaries are correlated to those in the Instructor's Resource Manual

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