

## Identifying Reaction Types & Balancing Equations

<b>Step 1: Name the type of chemical reaction</b>	<b>Step 2: Balance the following chemical equations:</b>
Ex. Combination Reaction	<u>2</u> N <sub>2</sub> + <u>6</u> H <sub>2</sub> $\square$ <u>4</u> NH <sub>3</sub>
1.	<u>    </u> KClO <sub>3</sub> $\square$ <u>    </u> KCl + <u>    </u> O <sub>2</sub>
2.	<u>    </u> NaCl + <u>    </u> F <sub>2</sub> $\square$ <u>    </u> NaF + <u>    </u> Cl <sub>2</sub>
3.	<u>    </u> H <sub>2</sub> + <u>    </u> O <sub>2</sub> $\square$ <u>    </u> H <sub>2</sub> O
4.	<u>    </u> AgNO <sub>3</sub> + <u>    </u> MgCl <sub>2</sub> $\square$ <u>    </u> AgCl + <u>    </u> Mg(NO <sub>3</sub> ) <sub>2</sub>
5.	<u>    </u> AlBr <sub>3</sub> + <u>    </u> K <sub>2</sub> SO <sub>4</sub> $\square$ <u>    </u> KBr + <u>    </u> Al <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>
6.	<u>    </u> CH <sub>4</sub> + <u>    </u> O <sub>2</sub> $\square$ <u>    </u> CO <sub>2</sub> + <u>    </u> H <sub>2</sub> O
7.	<u>    </u> C <sub>3</sub> H <sub>8</sub> + <u>    </u> O <sub>2</sub> $\square$ <u>    </u> CO <sub>2</sub> + <u>    </u> H <sub>2</sub> O
8.	<u>    </u> FeCl <sub>3</sub> + <u>    </u> NaOH $\square$ <u>    </u> Fe(OH) <sub>3</sub> + <u>    </u> NaCl
9.	<u>    </u> P + <u>    </u> O <sub>2</sub> $\square$ <u>    </u> P <sub>2</sub> O <sub>5</sub>
10.	<u>    </u> Na + <u>    </u> H <sub>2</sub> O $\square$ <u>    </u> NaOH + <u>    </u> H <sub>2</sub>
11.	<u>    </u> Ag <sub>2</sub> O $\square$ <u>    </u> Ag + <u>    </u> O <sub>2</sub>
12.	<u>    </u> CO <sub>2</sub> + <u>    </u> H <sub>2</sub> O $\square$ <u>    </u> C <sub>6</sub> H <sub>12</sub> O <sub>6</sub> + <u>    </u> O <sub>2</sub>
13.	<u>    </u> K + <u>    </u> MgBr <sub>2</sub> $\square$ <u>    </u> KBr + <u>    </u> Mg
14.	<u>    </u> HNO <sub>3</sub> (aq) + <u>    </u> Ba(OH) <sub>2</sub> (aq) $\square$ <u>    </u> Ba(NO <sub>3</sub> ) <sub>2</sub> (aq) + <u>    </u> H <sub>2</sub> O (l)
15.	<u>    </u> C <sub>3</sub> H <sub>8</sub> (g) + <u>    </u> O <sub>2</sub> (g) $\square$ <u>    </u> CO <sub>2</sub> (g) + <u>    </u> H <sub>2</sub> O(g)
16.	<u>    </u> Al(s) + <u>    </u> Fe <sub>2</sub> O <sub>3</sub> (aq) $\square$ <u>    </u> Al <sub>2</sub> O <sub>3</sub> (aq) + <u>    </u> Fe(s)
17.	<u>    </u> Al(s) + <u>    </u> O <sub>2</sub> (g) $\square$ <u>    </u> Al <sub>2</sub> O <sub>3</sub> (s)
18.	<u>    </u> P <sub>4</sub> + <u>    </u> O <sub>2</sub> $\square$ 2 P <sub>2</sub> O <sub>5</sub>
19.	<u>2</u> CaCO <sub>3</sub> $\square$ <u>    </u> CaO + <u>    </u> CO <sub>2</sub>
20.	<u>    </u> H <sub>2</sub> + <u>    </u> O <sub>2</sub> $\square$ 2H <sub>2</sub> O

# Balancing Equation Reaction Types Answer Key

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