Name: \_\_\_\_\_

Date: \_\_\_\_\_

## **Balancing Chemical Equations**

Balance the following chemical equations.

1. 
$$\underline{2}$$
 Fe +  $\underline{3}$  H<sub>2</sub>SO<sub>4</sub>  $\longrightarrow$   $\underline{1}$  Fe<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> +  $\underline{3}$  H<sub>2</sub>

2. 
$$\underline{1}$$
 CH<sub>4</sub> +  $\underline{2}$  O<sub>2</sub>  $\longrightarrow$   $\underline{1}$  CO<sub>2</sub> +  $\underline{2}$  H<sub>2</sub>O

3. 
$$1$$
 SiCl<sub>4</sub>( $\ell$ ) +  $2$  H<sub>2</sub>O( $\ell$ )  $\longrightarrow$   $1$  SiO<sub>2</sub>( $s$ ) +  $4$  HCl(aq)

4. 
$$\underline{2}$$
 AgI +  $\underline{1}$  Na<sub>2</sub>S  $\longrightarrow$   $\underline{1}$  Ag<sub>2</sub>S +  $\underline{2}$  NaI

5. 
$$\underline{4}$$
 NH<sub>3</sub> +  $\underline{5}$  O<sub>2</sub>  $\longrightarrow$   $\underline{4}$  NO +  $\underline{6}$  H<sub>2</sub>O

6. 1 FeO<sub>3</sub>(s) + 3 CO(g) 
$$\longrightarrow$$
 1 Fe( $\ell$ ) + 3 CO<sub>2</sub>(g)

7. 
$$\underline{1}$$
 SiO<sub>2</sub> +  $\underline{4}$  HF  $\longrightarrow$   $\underline{1}$  SiF<sub>4</sub> +  $\underline{2}$  H<sub>2</sub>O

9. 
$$4 (NH_4)_3PO_4 + 3 Pb(NO_3)_4 \longrightarrow 1 Pb_3(PO_4)_4 + 12 NH_4NO_3$$

# **Balancing Chemical Equations Unit 4 Answer Key**

Spencer L. Seager, Michael R. Slabaugh

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Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998

Prentice Hall Chemistry, 2000

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