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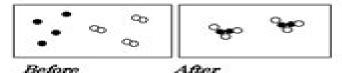
# Chemistry - Balancing Equations

Balance the following equations by inserting the proper coefficients. For selected reactions, draw *Before* and *After* particle diagrams to show the particles involved in the reaction. Be sure to provide a key.

10.3

#18

- 1.  $\underline{1}$  C +  $\underline{1}$  H<sub>2</sub>O  $\rightarrow$   $\underline{1}$  CO +  $\underline{1}$  H<sub>2</sub>
- 2.  $2 \text{ MgO} \rightarrow 2 \text{ Mg} + 1 \text{ O}_2$
- 3. \_4\_Al + \_3\_O<sub>2</sub> → \_2\_Al<sub>2</sub>O<sub>3</sub>



- 4.  $1 \text{ Zn} + 1 \text{ H}_2\text{SO}_4 \rightarrow 1 \text{ ZnSO}_4 + 1 \text{ H}_2$
- 5.  $1 \text{ Cl}_2 + 2 \text{ KI} \rightarrow 2 \text{ KCl} + 1 \text{ I}_2$
- 6. 2 CuCl  $\rightarrow$  2 Cu + 1 Cl<sub>2</sub>
- 2 Na + 1 Cl<sub>2</sub> → 2 NaCl
- 8.  $2\text{Al} + 6\text{HCl} \rightarrow 2\text{AlCl}_3 + 3\text{H}_2$



- 9.  $2 \text{ Fe}_2\text{O}_3 \rightarrow 4 \text{ Fe} + 3 \text{ O}_2$
- 10.  $\underline{4} P + \underline{5} O_2 \rightarrow \underline{2} P_2 O_5$
- 11.  $\underline{1}$  Mg +  $\underline{2}$  HCl  $\rightarrow$   $\underline{1}$  MgCl<sub>2</sub> +  $\underline{1}$  H<sub>2</sub>
- 12.  $3 \text{ H}_2 + 1 \text{ N}_2 \rightarrow 2 \text{ NH}_3$
- 13.  $\underline{1}$  BaCl<sub>2</sub> +  $\underline{1}$  H<sub>2</sub>SO<sub>4</sub>  $\rightarrow$   $\underline{1}$  BaSO<sub>4</sub> +  $\underline{2}$  HCl
- 14. <u>1 CH4 + 2 O2 → 1 CO2 + 2 H2O</u>



#14 Before After

es Plea

 $\phi^{O}$ o

# **Balancing Equations About Chemistry Answers**

**Chris McMullen** 

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