

# Balancing Equations Worksheet

*Double replacement reaction*

- 1)  $\underline{\hspace{1cm}} \text{Na}_3\text{PO}_4 + \underline{\hspace{1cm}} \text{KOH} \rightarrow \underline{\hspace{1cm}} \text{NaOH} + \underline{\hspace{1cm}} \text{K}_3\text{PO}_4$  *Na PO4*
- 2)  $\underline{\hspace{1cm}} \text{MgF}_2 + \underline{\hspace{1cm}} \text{Li}_2\text{CO}_3 \rightarrow \underline{\hspace{1cm}} \text{MgCO}_3 + \underline{\hspace{1cm}} \text{LiF}$
- 3)  $\underline{\hspace{1cm}} \text{P}_4 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{P}_2\text{O}_3$
- 4)  $\underline{\hspace{1cm}} \text{RbNO}_3 + \underline{\hspace{1cm}} \text{BeF}_2 \rightarrow \underline{\hspace{1cm}} \text{Be}(\text{NO}_3)_2 + \underline{\hspace{1cm}} \text{RbF}$
- 5)  $\underline{\hspace{1cm}} \text{AgNO}_3 + \underline{\hspace{1cm}} \text{Cu} \rightarrow \underline{\hspace{1cm}} \text{Cu}(\text{NO}_3)_2 + \underline{\hspace{1cm}} \text{Ag}$
- 6)  $\underline{\hspace{1cm}} \text{CF}_4 + \underline{\hspace{1cm}} \text{Br}_2 \rightarrow \underline{\hspace{1cm}} \text{CBr}_4 + \underline{\hspace{1cm}} \text{F}_2$
- 7)  $\underline{\hspace{1cm}} \text{HCN} + \underline{\hspace{1cm}} \text{CuSO}_4 \rightarrow \underline{\hspace{1cm}} \text{H}_2\text{SO}_4 + \underline{\hspace{1cm}} \text{Cu}(\text{CN})_2$
- 8)  $\underline{\hspace{1cm}} \text{GaF}_3 + \underline{\hspace{1cm}} \text{Cs} \rightarrow \underline{\hspace{1cm}} \text{CsF} + \underline{\hspace{1cm}} \text{Ga}$
- 9)  $\underline{\hspace{1cm}} \text{BaS} + \underline{\hspace{1cm}} \text{PtF}_2 \rightarrow \underline{\hspace{1cm}} \text{BaF}_2 + \underline{\hspace{1cm}} \text{PtS}$
- 10)  $\underline{\hspace{1cm}} \text{N}_2 + \underline{\hspace{1cm}} \text{H}_2 \rightarrow \underline{\hspace{1cm}} \text{NH}_3$
- 11)  $\underline{\hspace{1cm}} \text{NaF} + \underline{\hspace{1cm}} \text{Br}_2 \rightarrow \underline{\hspace{1cm}} \text{NaBr} + \underline{\hspace{1cm}} \text{F}_2$
- 12)  $\underline{\hspace{1cm}} \text{Pb}(\text{OH})_2 + \underline{\hspace{1cm}} \text{HCl} \rightarrow \underline{\hspace{1cm}} \text{H}_2\text{O} + \underline{\hspace{1cm}} \text{PbCl}_2$
- 13)  $\underline{\hspace{1cm}} \text{AlBr}_3 + \underline{\hspace{1cm}} \text{K}_2\text{SO}_4 \rightarrow \underline{\hspace{1cm}} \text{KBr} + \underline{\hspace{1cm}} \text{Al}_2(\text{SO}_4)_3$
- 14)  $\underline{\hspace{1cm}} \text{CH}_4 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$

# Balancing Equations Worksheet Answer Key Na<sub>3</sub>PO<sub>4</sub>

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