

# BALANCING CHEMICAL EQUATIONS

Name \_\_\_\_\_

Rewrite and balance the equations below.

1.  $\text{N}_2 + \text{H}_2 \rightarrow \text{NH}_3$  \_\_\_\_\_
2.  $\text{KClO}_3 \rightarrow \text{KCl} + \text{O}_2$  \_\_\_\_\_
3.  $\text{NaCl} + \text{F}_2 \rightarrow \text{NaF} + \text{Cl}_2$  \_\_\_\_\_
4.  $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$  \_\_\_\_\_
5.  $\text{AgNO}_3 + \text{MgCl}_2 \rightarrow \text{AgCl} + \text{Mg}(\text{NO}_3)_2$  \_\_\_\_\_
6.  $\text{AlBr}_3 + \text{K}_2\text{SO}_4 \rightarrow \text{KBr} + \text{Al}_2(\text{SO}_4)_3$  \_\_\_\_\_
7.  $\text{CH}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$  \_\_\_\_\_
8.  $\text{C}_3\text{H}_8 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$  \_\_\_\_\_
9.  $\text{C}_8\text{H}_{18} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$  \_\_\_\_\_
10.  $\text{FeCl}_3 + \text{NaOH} \rightarrow \text{Fe}(\text{OH})_3 + \text{NaCl}$  \_\_\_\_\_
11.  $\text{P} + \text{O}_2 \rightarrow \text{P}_2\text{O}_5$  \_\_\_\_\_
12.  $\text{Na} + \text{H}_2\text{O} \rightarrow \text{NaOH} + \text{H}_2$  \_\_\_\_\_
13.  $\text{Ag}_2\text{O} \rightarrow \text{Ag} + \text{O}_2$  \_\_\_\_\_
14.  $\text{S}_8 + \text{O}_2 \rightarrow \text{SO}_3$  \_\_\_\_\_
15.  $\text{CO}_2 + \text{H}_2\text{O} \rightarrow \text{C}_6\text{H}_{12}\text{O}_6 + \text{O}_2$  \_\_\_\_\_
16.  $\text{K} + \text{MgBr}_2 \rightarrow \text{KBr} + \text{Mg}$  \_\_\_\_\_
17.  $\text{HCl} + \text{CaCO}_3 \rightarrow \text{CaCl}_2 + \text{H}_2\text{O} + \text{CO}_2$  \_\_\_\_\_

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## If8767

**KJ Lindholm-Leary**



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