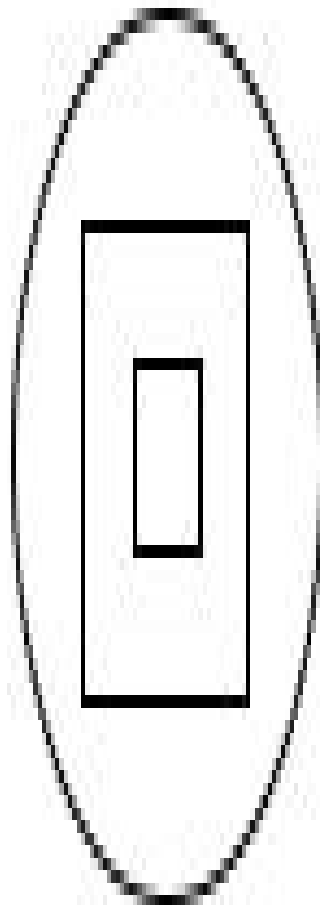
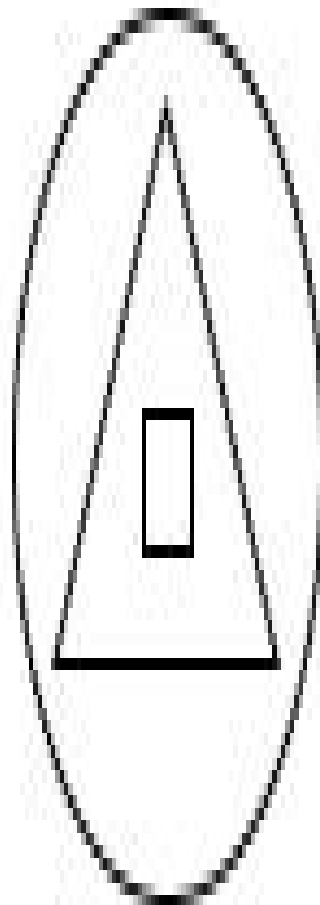


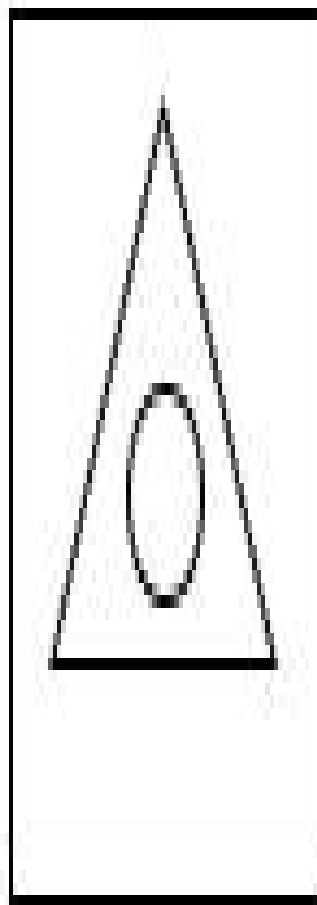
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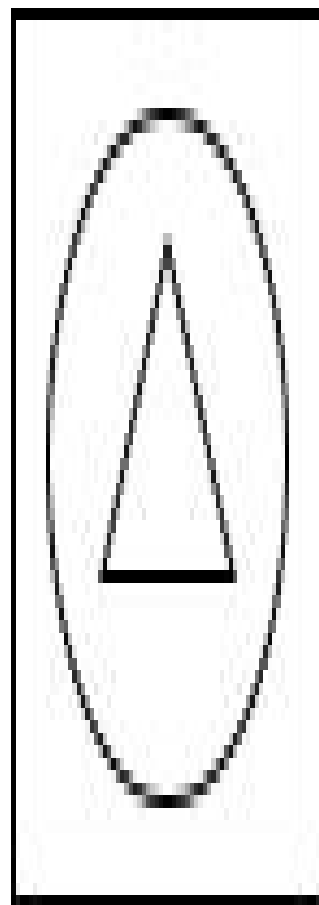
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U.S. Government Research & Development Reports, 1965 **Journal of applied psychology**, 1952 The Managerial Competencies of Twelve Corporate Librarians Marcy Murphy, 1988 Labor Arbitration Information System, 1989 *Psychological Testing and Assessment* Lewis R. Aiken, 1991 **Journal of Personnel Research**, 1923 Includes section Book reviews **Psychological Bulletin**, 1928 Vol 49 no 4 pt 2 July 1952 is the association's Publication manual *Bibliography of Scientific and Industrial Reports*, 1965-07 *Sales Management*, 1945 **Corporate Practice Series**, 1978 **Readings on Behavior in Organizations** Edgar F. Huse, James L. Bowditch, Dalmar Fisher, 1975 **A Fair Day's Output from a Fair Day's Work** Canada. Wartime Prices and Trade Board, 1945

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web goals understand that parts of an acid base titration be able to determine the K_a or K_b from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate K_{a1} and K_{a2} for a polyprotic acid by the end of this lab students should be able to

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 $\text{HCl aq} + \text{NaOH aq} \rightarrow \text{NaCl aq} + \text{H}_2\text{O l}$
 1 mole of HCl would be fully neutralized by one mole of NaOH

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web conclusion the identification of an unknown solid acid the purpose of this experiment was to determine the standardization of sodium hydroxide NaOH and oxalic acid $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ and the molar mass of an unknown solid acid using the molar mass of the standardized sodium hydroxide solution

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web this point of titration is referred to as the equivalence point considering the 1 1 stoichiometry of this acid base reaction
 $\text{NaOH (aq)} + \text{C}_6\text{H}_4\text{COOH (s)} \rightarrow \text{C}_6\text{H}_4\text{COONa (aq)} + \text{H}_2\text{O (l)}$
the point of equivalence is the point of titration when the number of moles of naoh added is equal to the number of moles of khp in the