

# CONJUGATE ACID-BASE PAIRS

Name \_\_\_\_\_

In the exercise, Bronsted-Lowry Acids and Bases, it was shown that after an acid has given up its proton, it is capable of getting back that proton and acting as a base. Conjugate base is what is left after an acid gives up a proton. The stronger the acid, the weaker the conjugate base. The weaker the acid, the stronger the conjugate base.

Fill in the blanks in the table below.

Conjugate Pairs

	ACID	BASE	EQUATION
1.	$\text{H}_2\text{SO}_4$	$\text{HSO}_4^-$	$\text{H}_2\text{SO}_4 \leftrightarrow \text{H}^+ + \text{HSO}_4^-$
2.	$\text{H}_3\text{PO}_4$		
3.		$\text{F}^-$	
4.		$\text{NO}_3^-$	
5.	$\text{H}_2\text{PO}_4^-$		
6.	$\text{H}_2\text{O}$		
7.		$\text{SO}_4^{2-}$	
8.	$\text{HPO}_4^{2-}$		
9.	$\text{NH}_4^+$		
10.		$\text{H}_2\text{O}$	

Which is a stronger base,  $\text{HSO}_4^-$  or  $\text{H}_2\text{PO}_4^-$ ? \_\_\_\_\_

Which is a weaker base,  $\text{Cl}^-$  or  $\text{NO}_3^-$ ? \_\_\_\_\_

# Conjugate Acid Base Pairs Answer Key Chemistry 1 7

**Ronald Percy Bell**



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