# Stoichiometry "End of Chapter" IN CLASS Practice Assessment

thackground into -

Why are the coefficients in a balanced equation necessary for dischiometry?
 Coefficients are necessary for establishing a male male ratio to convert between modes of different substances.

Skills - Using a balanced equation, calculate using mole to mole ratics & Using a balanced equation, given the grams of reactant, calc. the grams of product

2. Consider the following unbalanced equation:

4 HCl + MnO, -> 2 H<sub>2</sub>O + MnCl<sub>2</sub> + Cl<sub>2</sub>

There are 5.00g of manganese(N) oxide present, calculate the grame of chiorine gas. 5.00g MnO<sub>2</sub> × 1mol MnO<sub>2</sub> × 1mol Cl<sub>2</sub> × 70.9gCl<sub>3</sub> = 4.00g Cl<sub>3</sub> = 600 g Cl<sub>3</sub> = 600

Skill - After calculating theoretical yield in grams, calculate the percent yield.

You are assigned to complete a lab. Before the reaction is complete, you calculate that a
reaction should produce 4 0g Culpi. After the lab is over you find that 2.5g Culpi was
collected. What would the percent yield be for that reaction completed within the lab?
No yield — Lactual/Theoretical() x 2.00
 — C2.0ex/L0ex) x 1.00 — 200s.

Skill - Given experimental data, determine the limiting reactant and the excess reactant

4. Calcium cartionate continues with gaseous hydrogen chloride to produce solid calcium chloride, gaseous carbon disolde, and water vapor. What is the percent yield if 155g of calcium carbonate combines with 250 0g of hydrogen chloride and only 342 0g of calcium chloride in actually obtained within the list setting? Complete and show work for all 5 steps of a stotchomotry problem.

CaCO<sub>3</sub> + 2 HCl → CaCl<sub>3</sub> + CO<sub>3</sub> + H<sub>2</sub>O

130<sub>9</sub> 142,9 actual

155g CaCO, x 1mol CaCO<sub>2</sub>/100.09g CaCO<sub>2</sub> ~ 1.55mol CaCO<sub>2</sub> /1mol in non ~ 1.55 nons 250a HCL x 1mol HCL36.458g HCl ~ 6.057 mol HCl /2mol in non ~ 3.43 runs

1.55mol CaCO,  $\times$  (1mol CaCly1mol CaCO)  $\times$  (110.90gCaCly1mol CaCly) = 172.02n

%yield = [actual/theoretical] x 100 = [142.0g/172.02g] x 100 = 82.55%

# **Chapter Review Stoichiometry Section Answer Key**

Mary Millhollon, Richard H. Langley

#### **Chapter Review Stoichiometry Section Answer Key:**

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