

## Stoichiometry WorkSheet #1: Worked Solutions

Answer the following questions on your own paper. Show all work. Circle the final answer, giving units and the correct number of significant figures.

1. Based on the following equation, how many moles of each product are produced when 5.9 moles of  $\text{Zn}(\text{OH})_2$  are reacted with  $\text{H}_3\text{PO}_4$ ? (You need to balance the equation.)



5.9 moles $\text{Zn}(\text{OH})_2$	1 mole $\text{Zn}_3(\text{PO}_4)_2$ 3 moles $\text{Zn}(\text{OH})_2$	= 2.0 moles $\text{Zn}_3(\text{PO}_4)_2$
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5.9 moles $\text{Zn}(\text{OH})_2$	6 moles $\text{H}_2\text{O}$ 3 moles $\text{Zn}(\text{OH})_2$	= 12 moles $\text{H}_2\text{O}$
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# Chapter Stoichiometry Exercises Answers

**Morris Hein, Susan Arena, Cary Willard**



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