1. How many moles of sodium will react with water to produce 4.0 mol of hydrogen in the following reaction?

$$2 \text{ Na (s)} + 2 \text{ H}_2\text{O (l)} \rightarrow 2 \text{ NaOH (aq)} + \text{H}_2 \text{ (g)}$$

Ammonia is made industrially by reacting nitrogen and hydrogen under high pressure and temperature and in the presence of a catalyst. The equation is:

$$3 H_2(g) + N_2(g) \rightarrow 2 NH_3(g)$$

If 4.0 mol of H₂ react, how many moles of NH₃ will be produced?

3. Phosphorous burns in oxygen to produce phosphorous oxide in the following reaction:

$$4 P (s) + 5 O_2 (g) \rightarrow P_4O_{10} (s)$$

a. What amount (in moles) of phosphorous will be needed to produce 3.25 mol of P_4O_{10} ?

b. How many moles of P₄O₁₀ are produced from 2.74 moles of phosphorous?

4. How many moles of lithium chloride will be formed by the reaction of chlorine with 0.046 mol of lithium bromide in the following reaction?

$$2 \text{ LiBr (aq)} + \text{Cl}_2(g) \rightarrow 2 \text{ LiCl (aq.)} + \text{Br}_2(g)$$

5. Aluminum will react with sulfuric acid in the following reaction.

$$2 \text{ Al (s)} + 3 \text{ H}_2 \text{SO}_4 \text{ (aq)} \rightarrow \text{Al}_2 (\text{SO}_4)_3 + 3 \text{ H}_2 \text{ (g)}$$

How many moles of H₂SO₄ will react with 18 mol of Al?

Chapter Stoichiometry Worksheet Answers

Jianjun Gao

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