

## Mixed Gas Laws Worksheet - Solutions

- 1) How many moles of gas occupy 98 L at a pressure of 2.8 atmospheres and a temperature of 292 K?

$$n = \frac{PV}{RT} = \frac{(2.8 \text{ atm})(98 \text{ L})}{(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(292 \text{ K})} = 11 \text{ moles of gas}$$

- 2) If 5.0 moles of  $\text{O}_2$  and 3.0 moles of  $\text{N}_2$  are placed in a 30.0 L tank at a temperature of  $25^\circ\text{C}$ , what will the pressure of the resulting mixture of gases be?  $25^\circ\text{C} = 298 \text{ K}$

$$\text{O}_2: P = \frac{nRT}{V} = \frac{(5.0 \text{ mol})(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(298 \text{ K})}{(30.0 \text{ L})} = 4.1 \text{ atm}$$

$$\text{N}_2: P = \frac{nRT}{V} = \frac{(3.0 \text{ mol})(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(298 \text{ K})}{(30.0 \text{ L})} = 2.4 \text{ atm}$$

$$P_{\text{Tot}} = P_{\text{O}_2} + P_{\text{N}_2} = 4.1 \text{ atm} + 2.4 \text{ atm} = 6.5 \text{ atm}$$

- 3) A balloon is filled with 35.0 L of helium in the morning when the temperature is  $20.0^\circ\text{C}$ . By noon the temperature has risen to  $45.0^\circ\text{C}$ . What is the new volume of the balloon?

$$T_1 = 20.0^\circ\text{C} = 293 \text{ K}, V_1 = 35.0 \text{ L}, T_2 = 45.0^\circ\text{C} = 318 \text{ K}, V_2 = ?$$

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(35.0 \text{ L})(318 \text{ K})}{(293 \text{ K})} = 38.0 \text{ L}$$

- 4) A 35 L tank of oxygen is at 315 K with an internal pressure of 190 atmospheres. How many moles of gas does the tank contain?

$$n = \frac{PV}{RT} = \frac{(190 \text{ atm})(35 \text{ L})}{(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(315 \text{ K})} = 260 \text{ moles of gas}$$

- 5) A balloon that can hold 85 L of air is inflated with 3.5 moles of gas at a pressure of 1.0 atmosphere. What is the temperature in  $^\circ\text{C}$  of the balloon?

$$T = \frac{PV}{nR} = \frac{(1 \text{ atm})(85 \text{ L})}{(3.5 \text{ mol})(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})} = 296 \text{ K} = 23^\circ\text{C}$$

- 6)  $\text{CaCO}_3$  decomposes at  $1200^\circ\text{C}$  to form  $\text{CO}_2$  gas and  $\text{CaO}$ . If 25 L of  $\text{CO}_2$  are collected at  $1200^\circ\text{C}$ , what will the volume of this gas be after it cools to  $25^\circ\text{C}$ ?

$$T_1 = 1200^\circ\text{C} = 1473 \text{ K}, V_1 = 25 \text{ L}, T_2 = 25^\circ\text{C} = 298 \text{ K}, V_2 = ?$$

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(25 \text{ L})(298 \text{ K})}{(1473 \text{ K})} = 5.1 \text{ L}$$

- 7) A helium balloon with an internal pressure of 1.00 atm and a volume of 4.50 L at  $20.0^\circ\text{C}$  is released. What volume will the balloon occupy at an altitude where the pressure is 0.600 atm and the temperature is  $-20.0^\circ\text{C}$ ?

$$P_1 = 1.00 \text{ atm}, V_1 = 4.50 \text{ L}, T_1 = 20.0^\circ\text{C} = 293 \text{ K}, P_2 = 0.600 \text{ atm}, V_2 = ?, T_2 = -20.0^\circ\text{C} = 253 \text{ K}$$

$$V_2 = \frac{P_1 V_1 T_2}{T_1 P_2} = \frac{(1.00 \text{ atm})(4.50 \text{ L})(253 \text{ K})}{(293 \text{ K})(0.600 \text{ atm})} = 6.48 \text{ L}$$

- 8) There are 135 L of gas in a container at a temperature of  $260^\circ\text{C}$ . If the gas was cooled until the volume decreased to 75 L, what would the temperature of the gas be?

$$T_1 = 260^\circ\text{C} = 533 \text{ K}, V_1 = 135 \text{ L}, T_2 = ?, V_2 = 75 \text{ L}$$

$$T_2 = \frac{V_2 T_1}{V_1} = \frac{(75 \text{ L})(533 \text{ K})}{(135 \text{ L})} = 296 \text{ K} = 23^\circ\text{C}$$

# Chapter 14.5 Mixed Gas Laws Problems Answers

**J Ma**



## Chapter 14 5 Mixed Gas Laws Problems Answers:

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Cutnell, Kenneth W. Johnson, David Young, Shane Stadler, 2021-10-05 In the newly revised Twelfth Edition of Physics Volume 1 an accomplished team of physicists and educators delivers an accessible and rigorous approach to the skills students need to succeed in physics education Readers will learn to understand foundational physics concepts solve common physics problems and see real world applications of the included concepts to assist in retention and learning The text includes Check Your Understanding questions Math Skills boxes multi concept problems and worked examples The first volume of a two volume set Volume 1 explores ideas and concepts like Newton s Laws of Motion the Ideal Gas Law and kinetic theory Throughout students knowledge is tested with concept and calculation problems and team exercises that focus on cooperation and learning Technical Physics Frederick Bueche, David L. Wallach, 1994-03-31 The basic nature of the text remains the same but numerous modifications have been made to enhance its teachability Photometry units and definitions are now consistent with the latest publications A discussion of ideal fluid flow with Bernoulli s equation and the conservation of mass has been added An expanded mathematics section of the appendix includes logarithms and their equations as well as approximations and trigonometric identities Problems are greater in both number and variety **Challenging Problems for Physics** Boris Korsunsky, Raymond A. Serway, 1995 Study Guide for Atoms, Molecules, and Life Michael S. Matta, 1981 An Introduction to Thermodynamics, for Engineering Students John Mills, 1910 *Applied Mechanics Reviews* , 1993

**Physics** Eugene Hecht, 2003 Student text An Introduction to Physics Measurement The Language of Physics Kinematics Speed Velocity Speed Velocity Relative Motion Kinematics Acceleration The Concept of Acceleration Uniformly Accelerated Motion Free Fall Newton s Three Laws The Three Laws Dynamics Statics Centripetal Force Gravity Centripetal Force Gravity The Cosmic Force Energy The Transfer of Energy Mechanical Energy Conservation of Mechanical Energy Momentum Collisions Linear Momentum Rotational Motion The Kinematics of Rotation Rotational Equilibrium The Dynamics of Rotation Solids Liquids Gases Atoms Matter Fluid Statics Fluid Dynamics Elasticity Oscillations Elasticity Harmonic Motion Waves Sound Mechanical Waves Sound Thermal Properties of Matter Temperature Thermal Expansion The Gas Laws Heat Thermal Energy Thermal Energy Change of State The Transfer of Thermal Energy Thermodynamics The First Law of Thermodynamics Cyclic Processes Engines Refrigerators The Second Law of Thermodynamics Electrostatics Forces Electromagnetic Charge The Electric Force The Electric Field Electrostatics Energy Electric Potential Capacitance Direct Current Flowing Electricity Resistance Circuits Circuit Principles Network Analysis Optional Magnetism Magnets the Magnetic Field Electrodynamics Magnetic Force Electromagnetic Induction Electromagnetically Induced emf Generators Self Induction AC Electronics Alternating Current R L C AC Networks Optional Electronics Optional Radiant Energy Light The Nature of Light The Electromagnetic Photon Spectrum The Propagation of Light Scattering Reflection Refraction The World of Color Geometrical Optics Instruments Lenses Mirrors Physical Optics Polarization Interference Diffraction Special Relativity Before the Special Theory The Special Theory of Relativity Relativistic Dynamics The Origins of Modern Physics Subatomic Particles

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## Table of Contents Chapter14 5 Mixed Gas Laws Problems Answers

1. Understanding the eBook Chapter14 5 Mixed Gas Laws Problems Answers
  - The Rise of Digital Reading Chapter14 5 Mixed Gas Laws Problems Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter14 5 Mixed Gas Laws Problems Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter14 5 Mixed Gas Laws Problems Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter14 5 Mixed Gas Laws Problems Answers
  - Personalized Recommendations
  - Chapter14 5 Mixed Gas Laws Problems Answers User Reviews and Ratings
  - Chapter14 5 Mixed Gas Laws Problems Answers and Bestseller Lists
5. Accessing Chapter14 5 Mixed Gas Laws Problems Answers Free and Paid eBooks
  - Chapter14 5 Mixed Gas Laws Problems Answers Public Domain eBooks
  - Chapter14 5 Mixed Gas Laws Problems Answers eBook Subscription Services
  - Chapter14 5 Mixed Gas Laws Problems Answers Budget-Friendly Options
6. Navigating Chapter14 5 Mixed Gas Laws Problems Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter14 5 Mixed Gas Laws Problems Answers Compatibility with Devices
  - Chapter14 5 Mixed Gas Laws Problems Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter14 5 Mixed Gas Laws Problems Answers
  - Highlighting and Note-Taking Chapter14 5 Mixed Gas Laws Problems Answers
  - Interactive Elements Chapter14 5 Mixed Gas Laws Problems Answers
8. Staying Engaged with Chapter14 5 Mixed Gas Laws Problems Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Chapter14 5 Mixed Gas Laws Problems Answers
- 9. Balancing eBooks and Physical Books Chapter14 5 Mixed Gas Laws Problems Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter14 5 Mixed Gas Laws Problems Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter14 5 Mixed Gas Laws Problems Answers
  - Setting Reading Goals Chapter14 5 Mixed Gas Laws Problems Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter14 5 Mixed Gas Laws Problems Answers
  - Fact-Checking eBook Content of Chapter14 5 Mixed Gas Laws Problems Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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