

## MODULE C: SOLIDS, LIQUIDS & GASES

### Kinetic Molecular Theory & Phases of Matter

- The different phases of matter can be described based on the Kinetic Molecular Theory

a) According to kinetic theory, an increase in temperature will cause the average kinetic energy of the particles in the solid, liquid, and gas to increase.

↑ temperature    ↑ average kinetic energy

b) When a solid is heated, the temperature is raised and the solid changes to a liquid at its melting point. If the heating is continued, the liquid will change into a gas at its boiling point.

c) The heat energy is changed into the equivalent amount of kinetic energy. Thus, the kinetic energy content of the particles increases in the order:

solid < liquid < gas

### - Kinetic Molecular Theory

- 1) Matter is made up of particles
- 2) The particles are always moving

### - Kinetic energy

→ energy of movement

### - Kinetic Molecular Theory for Gases

a) A gas consists of small particles (atoms or molecules) which have negligible volume ( $V$  of gas molecule  $\approx 0$ ). This implies that:

$$V_{\text{container}} = V_{\text{molecule}} + V_{\text{free space}} \approx V_{\text{free space}}$$

and that there is plenty of space for compression.

b) Gas particles are separated relatively far from one another.

c) Gas particles behave independently from one another. Thus, inter-particle forces of attraction are negligible.

d) The gas particles are in rapid and continuous motion.

e) In the gaseous state, the gas particles have vibrational, rotational, and translational motion.

1) **Vibrational motion** - a particle vibrates about a fixed point

2) **Rotational motion** - a particle spins about a fixed point

3) **Translational motion** - a particle moves from one place to another

f) Pressure of a gas comes from molecules colliding with the walls of the container. Such collisions are perfectly elastic.

g) Average kinetic energy depends only on the absolute temperature in Kelvin (K)

$$\text{Average Kinetic Energy} = \frac{3}{2} RT, \quad R \rightarrow \text{gas law constant } (0.08206 \frac{\text{L-atm}}{\text{K-mol}})$$

$T \rightarrow \text{temperature (in Kelvin)}$

# Chapter Review Part B Concept Review Answers

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**Focus on physical science** Charles H. Heimler, 1989      Interactions of Life ,      *Fundamentals of Microbiology* Jeffrey C. Pommerville, 2017-05-02 Pommerville s Fundamentals of Microbiology Eleventh Edition makes the difficult yet essential concepts of microbiology accessible and engaging for students initial introduction to this exciting science      *Scott Foresman-Addison Wesley Mathematics* , 2008

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Instructor's Manual to Accompany Computer Communications and Networking Technologies ,2002      *The Philosophical Review* Jacob Gould Schurman,James Edwin Creighton, Frank Thilly, Gustavus Watts Cunningham,1918 An international journal of general philosophy      Study Guide for Use with Macroeconomics [by] J. Bradford Delong Martha L. Olney,J. Bradford De Long,2002      **Real World Psychology** Catherine A. Sanderson,Karen R. Huffman,2019-12-24 Real World Psychology balances comprehensive coverage of the key concepts in introductory psychology with a concise presentation style and engages students with current and interesting research that explores these concepts in real life contexts Real World Psychology features the incomparable author team of Karen Huffman Palomar College and Catherine Sanderson Amherst College who create an outstanding text that is appealing to students and instructors at a wide range of academic institutions The new edition has been thoroughly updated and features a new focus on Scientific Thinking and Practical Applications underscoring the fact that connecting the principles of psychological science to everyday life is critical to student engagement and ultimately key to their success not only in the introductory psychology course but in whatever their chosen field of study and in everyday life Students will leave the course with an appreciation of how a basic yet scientific understanding of human behavior can benefit them in their studies in their personal lives and in their professional endeavors

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