

# Stoichiometry Review

**\*\*YOU MUST SHOW ALL WORK AND YOUR ANSWERS MUST INCLUDE THE PROPER NUMBER OF SIG FIGS AND COMPLETE UNITS IN ORDER TO RECEIVE CREDIT FOR THE PROBLEM.\*\***

BALANCE THE FOLLOWING EQUATIONS TO USE IN QUESTIONS 5 – 14:



PERFORM THE FOLLOWING STOICHIOMETRIC CALCULATIONS:

5. Zinc reacts with hydrochloric acid to produce zinc chloride and hydrogen. How many moles of HCl are required to produce 7.50 moles of ZnCl<sub>2</sub>?

$$\frac{7.50 \text{ mol ZnCl}_2}{1 \text{ mol ZnCl}_2} \times \frac{2 \text{ mol HCl}}{1 \text{ mol ZnCl}_2} = 15.0 \text{ mol HCl}$$

6. Copper metal reacts with silver nitrate to form silver and copper(II) nitrate. How many grams of copper are required to form 250 g of silver?

$$\frac{250 \text{ g Ag}}{107.868 \text{ g Ag}} \times \frac{1 \text{ mol Ag}}{1 \text{ mol Ag}} \times \frac{1 \text{ mol Cu}}{2 \text{ mol Ag}} \times \frac{63.546 \text{ g Cu}}{1 \text{ mol Cu}} = 74 \text{ g Cu}$$

7. When aluminum is burned in excess oxygen, aluminum oxide is produced. How many grams of oxygen are required to produce 0.75 moles of Al<sub>2</sub>O<sub>3</sub>?

$$\frac{0.75 \text{ mol Al}_2\text{O}_3}{1 \text{ mol Al}_2\text{O}_3} \times \frac{3 \text{ mol O}_2}{2 \text{ mol Al}_2\text{O}_3} \times \frac{31.998 \text{ g O}_2}{1 \text{ mol O}_2} = 36 \text{ g O}_2$$

8. How many grams of iron(III) chloride are produced when 15.3 g of iron react with excess chlorine gas?

$$\frac{15.3 \text{ g Fe}}{55.845 \text{ g Fe}} \times \frac{1 \text{ mol Fe}}{1 \text{ mol Fe}} \times \frac{2 \text{ mol FeCl}_3}{2 \text{ mol Fe}} \times \frac{162.204 \text{ g FeCl}_3}{1 \text{ mol FeCl}_3} = 44.4 \text{ g FeCl}_3$$

9. Copper metal reacts with silver nitrate to form silver and copper(II) nitrate. How many moles of silver will be produced from 3.65 moles of silver nitrate?

$$\frac{3.65 \text{ mol AgNO}_3}{2 \text{ mol AgNO}_3} \times \frac{2 \text{ mol Ag}}{1 \text{ mol AgNO}_3} = 3.65 \text{ mol Ag}$$

# Chapter Review Stoichiometry Key

**Morris Hein**



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Study Guide Darrell D. Ebbing, Susan M. Schelble, 1998     **Foundations of College Chemistry** Morris Hein, 1991. Features strong problem solving and carefully constructed problem sets.

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