

CHAPTER-1

SOLUTIONS

Topic- 1

Solutions and its Expression of Concentration, Solubility of Gases

Concepts Covered • Types of solution, • Molarity, • Normality, • ppm, • Mass by volume%, • Mole Fraction, • Volume by volume%, • Henry's Law



Revision Notes

- **Solution:** A homogeneous mixture of two or more pure substances is known as solution.
- **Two constituents of the solution are:**
 - (i) **Solute:** A substance that is dissolved in another substance in lesser amount, forming a solution. For example: Sugar, salt, etc.
 - (ii) **Solvent:** A substance in which another substance is dissolved in larger amount, forming a solution. For example: Water, milk, etc.
- Note:** Solvent determines the physical state of the solution.
- **Types of Solutions:** Any state of matter (solid, liquid or gas) can act both as a solvent and solute during the formation of a solution. Therefore, depending upon the physical states of solute and solvent, we can have following nine different types of solutions:

S. No.	Types of Solution	Solute	Solvent	Examples
1.	Solid – Solid	Solid	Solid	Alloys like brass, bronze, etc.
2.	Solid – Liquid	Solid	Liquid	Solution of sugar, salt, urea, etc., in water.
3.	Solid – Gas	Solid	Gas	Sublimation of substances like iodine, camphor etc. into air, dust or smoke particles in air.
4.	Liquid – Solid	Liquid	Solid	Hydrated salts, mercury in amalgamated zinc etc.
5.	Liquid – Liquid	Liquid	Liquid	Alcohol in water, benzene in toluene.
6.	Liquid – Gas	Liquid	Gas	Aerosol , water vapour in air.
7.	Gas – Solid	Gas	Solid	Hydrogen adsorbed in palladium.
8.	Gas – Liquid	Gas	Liquid	Aerated drinks.
9.	Gas – Gas	Gas	Gas	Mixture of gases, etc.



Key Word

Aerosol: It is suspension of fine solid or liquid particles in air, e.g.: Fog.

- **Aqueous solution:** A solution containing water as solvent is known as aqueous solution. For example: sugar solution.
- **Non-aqueous solution:** A solution containing solvent other than water is known as non-aqueous solution. For example: iodine dissolved in alcohol.
- **Saturated solution:** A solution in which no more solute can be dissolved at the same temperature is known as saturated solution.
- **Unsaturated solution:** A solution in which more amount of solute can be dissolved at the same temperature is known as unsaturated solution.
- **Method of expressing concentration of solution:** The concentration of solution is the amount of solute present in the given quantity of solute or solvent. It can be expressed in any of the following types:

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