

26. A system has reached equilibrium when:

- a) maximum entropy has been achieved
- b) minimum enthalpy has been achieved
- c) the rate of the forward reaction and reverse reaction is zero
- d) the concentrations of reactants and products have stopped changing

2. Your answer is d. Explain why Again, conc is a macroscopic property - stops changing once equlm is reached

27. Equilibrium is achieved when reactant and product concentrations are (equal/constant/zero)

constant

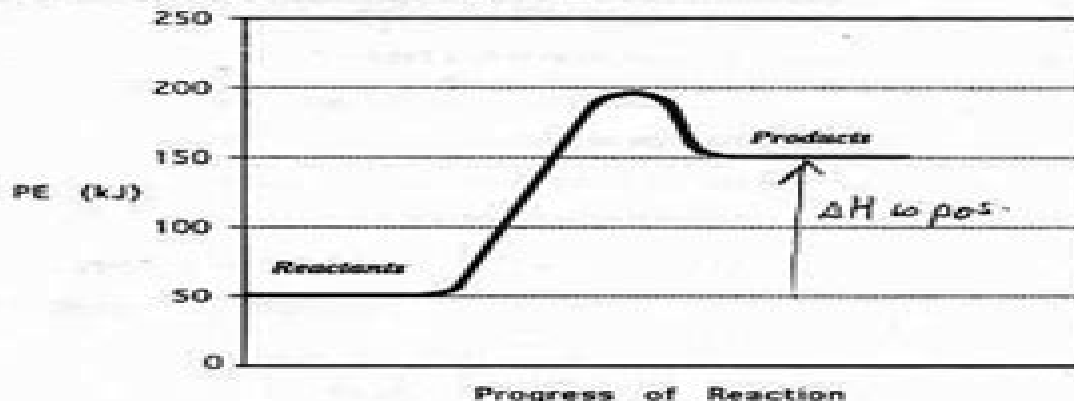
28. In a particular chemical reaction,  $\Delta H = +100 \text{ kJ}$ . When equilibrium has been established; it is found that a significant amount of product has formed, even though there is still some reactants left.

What has happened to entropy as this reaction was taking place? increased

Explain how you arrived at your answer rx. is endothermic ( $\Delta H$  pos.) so

tendency to min enthalpy favors reactants. In order for the rx. to occur at all, the tend. to max entropy must favor products (products have greater entropy)

29. Given the following potential energy diagram for a reaction:



Explain in terms of enthalpy and entropy, how you could end up with a fairly high ratio of products to reactants. endo ( $\Delta H$  pos.) tend. to min enthalpy favors reactants

If prod. to reactants ratio is high, tend. to max entropy must favor products. (products must have higher entropy)

# Chapter Study Guide Reaction Rates Answer Key

**Marcy Osgood, Karen A. Ocorr**



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