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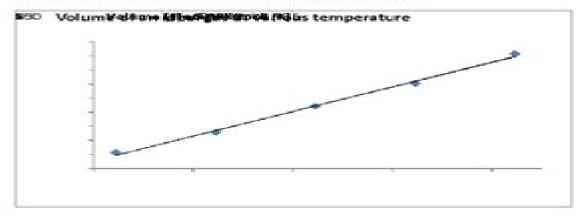
#### Chapter 6: Gases

Section A: Fill out this section and take notes as you watch this chapter's lecture video.

The average pressure at sea level is \_\_\_\_\_1 atm or \_\_\_\_\_760\_\_\_ mmHg or \_\_\_760\_\_ torr.

2. According to Boyle's Law, as the external pressure on a gas increases, it volume decreases.

Use the graph below to answer questions 3 and 4. The data on the graph is made up to represent what real data may look like from analyzing a real gas that is behaving nearly ideally. The volume of the gas is taken at various temperatures while holding the number of moles of gas and the pressure constant during the experiment.



3. Which of the following describes the relationship between volume and temperature in the graph above?

A) Volume is inversely proportional to temperature.

B) Volume is directly proportional to temperature.

C) Volume is proportional to the square root of temperature.

4. What does the slope represent in the graph above? Rearrange the ideal gas law to find this.

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