

12

STOICHIOMETRY

Practice Problems

In your notebook, solve the following problems.

SECTION 12.1 THE ARITHMETIC OF EQUATIONS

Use the 3-step problem-solving approach you learned in Chapter 1.

1. An apple pie needs 10 large apples, 2 cups (top and bottom), and 1 tablespoon of cinnamon. Write a balanced equation that fits this situation. How many apples are needed to make 20 pies?
2. Two moles of potassium chloride and three moles of oxygen are produced from the decomposition of two moles of potassium chlorate, $2\text{KClO}_3(\text{s})$. Write the balanced equation. How many moles of oxygen are produced from 12 moles of potassium chlorate?
3. Using the equation from problem 2, how many moles of oxygen are produced from 14 moles of potassium chlorate?
4. Two molecules of hydrogen react with one molecule of oxygen to produce two molecules of water. How many molecules of water are produced from 3.8×10^{23} molecules of oxygen? How many moles of water are produced from 22.5 moles of oxygen?

SECTION 12.2 CHEMICAL CALCULATIONS

1. Calculate the number of moles of hydrogen chloride produced from 10 moles of hydrogen.



2. Calculate the number of moles of chlorine needed to form 14 moles of iron(II) chloride.



3. Calculate the number of grams of nitrogen dioxide that are produced from 4 moles of nitric oxide.



4. Calculate the mass of oxygen produced from the decomposition of 75.0 g of potassium chlorate.



5. Calculate the mass of silver needed to react with chlorine to produce 94 g of silver chloride. Write a balanced equation first.

6. How many liters of carbon monoxide at STP are needed to react with 4.00 g of oxygen gas to produce carbon dioxide?



7. Calculate the number of liters of oxygen gas needed to produce 15.0 liters of dinitrogen monoxide. Assume all gases are at the same conditions of temperature and pressure.



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Steven S. Zumdahl, Paul B. Kelter



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studying an introductory chemistry course Teachers may also find this book a good resource for their lessons in stoichiometry **You Can Do Chemistry** Sunil Tanna,2018-11-19 A comprehensive guide to performing mole and stoichiometric calculations with numerous examples as well as questions and answers Covers calculations relating to solids solutions gases and electrolysis plus as limiting and excess reactants chemical yields atom economy and much more Fully up to date with the last international standards including the revised definition of mole which was agreed on November 16th 2018 **Chemistry Lesson Plans, Study Guides, and Lecture Notes** Doctor Chemistry,2017-08-13 The Ultimate Guide to Learning or Teaching Chemistry This book contains the real lecture notes and slide of a highly effective high school and college Chemistry teacher Teachers Never plan another lesson again Students Ace your upcoming exam This series covers all of the topics of High School Chemistry and General Chemistry including Accuracy and Significant Figures Mixtures Metric System Bonding Atomic Theory Periodic Table VSEPR Ionic and Covalent Bonding Geometric Bonding The Mole and Molar Mass Equation Balancing Thermodynamics Stoichiometry States of Matter Gas Laws and Calculations Reaction Calculations Acids and Bases Limiting Reagents Redox and Electro Chemistry Organic Chemistry Basics **Mole Concepts and Stoichiometry** Dharshi Bopegedera,2017

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Table of Contents Chapter Stoichiometry Guided Reading And Study Workbook Answers

1. Understanding the eBook Chapter Stoichiometry Guided Reading And Study Workbook Answers
 - The Rise of Digital Reading Chapter Stoichiometry Guided Reading And Study Workbook Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter Stoichiometry Guided Reading And Study Workbook Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter Stoichiometry Guided Reading And Study Workbook Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter Stoichiometry Guided Reading And Study Workbook Answers

- Personalized Recommendations
 - Chapter Stoichiometry Guided Reading And Study Workbook Answers User Reviews and Ratings
 - Chapter Stoichiometry Guided Reading And Study Workbook Answers and Bestseller Lists
5. Accessing Chapter Stoichiometry Guided Reading And Study Workbook Answers Free and Paid eBooks
 - Chapter Stoichiometry Guided Reading And Study Workbook Answers Public Domain eBooks
 - Chapter Stoichiometry Guided Reading And Study Workbook Answers eBook Subscription Services
 - Chapter Stoichiometry Guided Reading And Study Workbook Answers Budget-Friendly Options
 6. Navigating Chapter Stoichiometry Guided Reading And Study Workbook Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter Stoichiometry Guided Reading And Study Workbook Answers Compatibility with Devices
 - Chapter Stoichiometry Guided Reading And Study Workbook Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter Stoichiometry Guided Reading And Study Workbook Answers
 - Highlighting and Note-Taking Chapter Stoichiometry Guided Reading And Study Workbook Answers
 - Interactive Elements Chapter Stoichiometry Guided Reading And Study Workbook Answers
 8. Staying Engaged with Chapter Stoichiometry Guided Reading And Study Workbook Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter Stoichiometry Guided Reading And Study Workbook Answers
 9. Balancing eBooks and Physical Books Chapter Stoichiometry Guided Reading And Study Workbook Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter Stoichiometry Guided Reading And Study Workbook Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chapter Stoichiometry Guided Reading And Study Workbook Answers
 - Setting Reading Goals Chapter Stoichiometry Guided Reading And Study Workbook Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chapter Stoichiometry Guided Reading And Study Workbook Answers

- Fact-Checking eBook Content of Chapter Stoichiometry Guided Reading And Study Workbook Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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