

Chapter 18- Radioactivity and Nuclear Reactions

I. Radioactivity

- A. The nucleus of an atom is composed of protons and neutrons which comprise most of the atom's mass.
- B. The strong force causes protons and neutrons to be attracted to each other in the nucleus.
 - 1. The strong force is powerful only when neutrons and protons are closely packed together.
 - 2. The protons and neutrons in a large nucleus are held less tightly by the strong force than protons and neutrons in a small nucleus.
- C. Radioactivity—nuclear decay which happens when the strong force is not large enough to hold the nucleus together; the nucleus gives off matter and energy.
 - 1. Isotopes are atoms of the same element with varying numbers of neutrons.
 - 2. A nucleus with too many or too few neutrons compared to protons is radioactive.
 - 3. Atomic number—number of protons in an atom
 - 4. Mass number—number of protons and neutrons in a nucleus
- D. Uranium radioactivity was discovered in 1896 by Henri Becquerel; Marie and Pierre Curie discovered the radioactive elements polonium and radium in 1898.

II. Nuclear Decay

- A. Nuclear radiation—particles and energy are released from a decaying nucleus.
- B. Alpha particle—two protons and two neutrons with an electric charge of +2
 - 1. Alpha particles leave charged ions in their path when they travel through matter.
 - 2. Alpha particles are the least penetrating form of nuclear radiation.
 - 3. Alpha particles can cause serious biological damage.
- C. Beta particle—neutron decays into a proton and releases an electron at high speed; more penetrating than alpha particles
- D. Gamma ray—penetrating electromagnetic waves that carry energy, but have no mass or charge
- E. Transmutation—process of one element's changing to another through nuclear decay
- F. The half-life of a radioactive isotope is the length of time it takes half the nucleus to decay; half-lives vary from fractions of a second to billions of years.
- G. Carbon dating can be used to date once-living materials while uranium dating can be used to date rocks.

III. Detecting Radioactivity

- A. Radiation detectors are instruments used to identify ions formed when radiation passes through matter.
 - 1. A cloud chamber detects alpha or beta particles by means of a trail of condensed vapor.
 - 2. A bubble chamber detects radioactive particles by means of a bubble trail in a superheated liquid.
 - 3. Electroscopes can measure charged particles in the air.
- B. A Geiger counter measures radioactivity by producing an electric current when radiation is present.
- C. Background radiation comes from radon gas produced in the Earth's crust, from cosmic rays, and from radioactive isotopes in the body.

Chapter Review Radioactivity And Nuclear Reaction Answers

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Holt Physical Science Mapi M. Cuevas,1994 **Modern Chemistry** Holt Rinehart & Winston,Holt, Rinehart and Winston Staff,2001

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