

1. Ion-dipole force always require ____
 - a. an ion and a water molecule
 - b. a cation and a water molecule
 - c. an anion and a polar molecule
 - ✓ d. an ion and a polar molecule
 - e. a polar and a nonpolar molecule
2. Which substance in each of the following pairs would you expect to have a higher boiling point?
 - a. O_2 and Cl_2 **Cl_2**
 - b. SO_2 and CO_2 **SO_2**
 - c. HF and HI **HF**
 - d. F_2 and LiF **LiF**
3. The following compound have the same molecular forces (C_6H_{10}). Which one would you expect to have a higher boiling point? Explain your answer.



Molecule A has higher boiling point. Van der Waals forces are dependent on surface area. Both molecules have dispersion forces and have the same molecular mass but molecule A has larger surface area thus stronger IMF.

4. HCl (MW=36.5) has a dipole moment of 1.03 D and a boiling point of 190K. HBr (MW=80.9) has a dipole of 0.79 D and a boiling point of 206K. Which of the following statements is true?
 - a. HBr is more polar.
 - b. HCl has stronger intermolecular forces.
 - c. HCl has stronger London dispersion forces.
 - d. Both molecules have hydrogen bonding.
 - ✓ e. HBr has stronger intermolecular forces.
5. Hexane, C_6H_{14} (MW=86) has a boiling point of 68°C. Ethanol, CH_3CH_2OH (MW=46) has a boiling point of 78°C. Mark each of the following statements as TRUE or FALSE.
 - a. Ethanol must have stronger intermolecular attraction, based on its higher boiling point. **TRUE**
 - b. Ethanol has a higher boiling point because of greater London dispersion force. **FALSE**
 - c. Both hexane and ethanol have hydrogen bonding. **FALSE**
 - d. Ethanol has a higher boiling point due to hydrogen bonding. **TRUE**
 - e. Hydrogen bonding and London dispersion forces are at cross purposes here. (One favors ethanol, the other favors hexane.) In this case, the hydrogen bonding "wins". **TRUE**

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Crossley William C. Barlow



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