

CHAPTER 5 Electrons in Atoms

1. Orange light has a frequency of $4.8 \times 10^{14} \text{ s}^{-1}$. What is the energy of one quantum of orange light?
2. Which is greater, the energy of one photon of orange light or the energy of one quantum of radiation having a wavelength of $3.36 \times 10^{-8} \text{ m}$?
3. Use the relationships $E = h\nu$ and $c = \lambda\nu$ to write E in terms of h , c , and λ .
4. A radio station emits radiation at a wavelength of 2.90 m. What is the station's frequency in megahertz?
5. Record the frequency of your favorite radio station. What is the wavelength of the radiation emitted from the station?
6. List the sequence in which the following orbitals fill up: 1s, 2s, 3s, 4s, 5s, 6s, 7s, 2p, 3p, 4p, 5p, 6p, 7p, 3d, 4d, 5d, 6d, 4f, 5f.
7. Which element has the ground-state electron configuration $[\text{Kr}]5s^2 4d^{10} 5p^4$?
8. Which element has the ground-state electron configuration $[\text{Ar}]4s^2 3d^{10}$?
9. Write electron-dot structures for the following atoms.
 - a. $[\text{Ne}]3s^2 3p^3$
 - b. $[\text{Ar}]4s^2 3d^3$
 - c. potassium

10. Complete the following table.

Element	Symbol	Orbitals					Electron Configuration
		1s	2s	2p _x	2p _y	2p _z	
a. Nitrogen							$1s^2 2s^2 2p^3$
b.	F	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow\downarrow$	\uparrow	
c. Carbon							
d.							$1s^2 2s^1$

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