

Date: _____

Ionic and Covalent Dot Diagrams

Compound	Formula	Type of Bond	Dot Diagram
Chlorine	Cl_2	covalent	$14e^-$ $:\ddot{\text{Cl}}-\ddot{\text{Cl}}:$ linear
Silicon Dioxide	SiO_2	covalent	$\ddot{\text{O}}=\text{Si}=\ddot{\text{O}}$ linear
Barium Oxide	BaO	ionic	Before: $\text{Ba}^{+2} \quad :\ddot{\text{O}}^{2-}$ After: $[\text{Ba}]^{+2} \quad [:\ddot{\text{O}}:]^{2-}$
Aluminum Chloride	AlCl_3	ionic	Before: $\text{Al}^{+3} \quad 3 \times :\ddot{\text{Cl}}:^-$ After: $[\text{Al}]^{+3} \quad 3 \times [:\ddot{\text{Cl}}:]^{-1}$
Nitric Acid	HNO_3	covalent	$24e^-$ $\text{H}-\ddot{\text{O}}-\text{N}=\ddot{\text{O}}$ trigonal planar $:\ddot{\text{O}}:$

Date: _____

Ionic and Covalent Dot Diagrams

Compound	Formula	Type of Bond	Dot Diagram
Sodium Iodide	NaI	ionic	Before: $\text{Na}^{+1} \quad :\ddot{\text{I}}:^-$ After: $[\text{Na}]^{+1} \quad [:\ddot{\text{I}}:]^{-1}$
Hydrogen Sulfide	H_2S	covalent	$8e^-$ $\text{H}-\ddot{\text{S}}-\text{H}$ bent
Silicate Ion	SiO_3^{2-}	covalent	$24e^-$ $[\ddot{\text{O}}-\ddot{\text{Si}}-\ddot{\text{O}}]^{2-}$ pyramidal $:\ddot{\text{O}}:$
Calcium Fluoride	CaF_2	ionic	Before: $\text{Ca}^{+2} \quad 2 \times :\ddot{\text{F}}:^-$ After: $[\text{Ca}]^{+2} \quad 2 \times [:\ddot{\text{F}}:]^{-1}$
Hydrogen	H_2	covalent	$2e^-$ $\text{H}-\text{H}$ linear

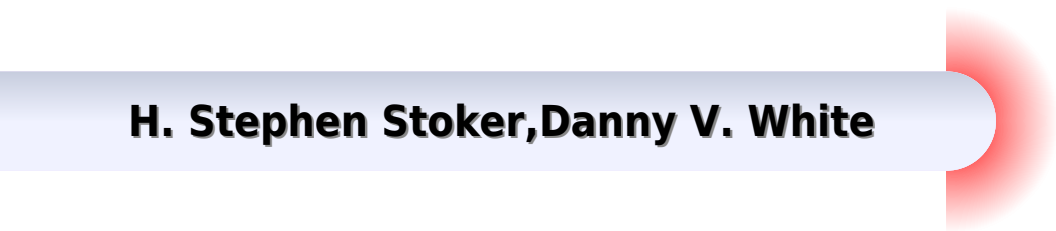
Draw the dot diagram for the following compounds.

Magnesium Phosphate	$\text{Mg}_3(\text{PO}_4)_2$	Ammonium Chloride	NH_4Cl
$3[\text{Mg}]^{+2} \quad [:\ddot{\text{O}}-\ddot{\text{P}}(\ddot{\text{O}})_3:]^{-3} \quad 2[\ddot{\text{O}}-\ddot{\text{P}}(\ddot{\text{O}})_3:]^{-3}$		$[\text{H}-\ddot{\text{N}}(\text{H})_3]^{+1} \quad [:\ddot{\text{Cl}}:]^{-1}$	
Chromium Cyanide	$\text{Cr}(\text{CN})_2$	Potassium Nitrate	KNO_3
$[\text{Cr}]^{+2} \quad 2[:\text{C}\equiv\text{N}:]^{-1}$		$[\text{K}]^{+1} \quad [:\ddot{\text{O}}-\text{N}(\ddot{\text{O}})_2:]^{-1}$	
Aluminum Phosphate	AlPO_4	Barium Sulfite	BaSO_3
$[\text{Al}]^{+3} \quad [:\ddot{\text{O}}-\ddot{\text{P}}(\ddot{\text{O}})_3:]^{-3}$		$[\text{Ba}]^{+2} \quad [:\ddot{\text{O}}-\ddot{\text{S}}(\ddot{\text{O}})_2:]^{-2}$	

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Chapter Test 20 Bonding And Chemical Reactions Answers

H. Stephen Stoker, Danny V. White



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