

12 • The Gas Laws**CHARLES'S LAW**

Charles' Law states the volume of a gas varies directly with the Kelvin temperature, assuming the pressure is constant. We use the following formulas:

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \quad \text{or} \quad V_1 \times T_2 = V_2 \times T_1$$

$$K = ^\circ C + 273$$

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Solve the following problems assuming a constant pressure. Assume all numbers are 3 significant figures.

1. A sample of nitrogen occupies a volume of 250 mL at $25^\circ C$. What volume will it occupy at $95^\circ C$?

$$\begin{array}{ccc} V_1 & T_1 & T_2 \\ 250 \text{ mL} & 298 \text{ K} & 368 \text{ K} \end{array}$$
2. Oxygen gas is at a temperature of $40^\circ C$ when it occupies a volume of 2.30 Liters. To what temperature should it be raised to occupy a volume of 6.50 Liters?

$$\begin{array}{ccc} T_1 & V_1 & V_2 \\ 313 \text{ K} & 2.30 \text{ L} & 6.50 \text{ L} \end{array}$$
3. Hydrogen gas was cooled from $150^\circ C$ to $50^\circ C$. Its new volume is 75.0 mL. What was its original volume?

$$\begin{array}{ccc} T_1 & T_2 & V_2 \\ 423 \text{ K} & 323 \text{ K} & 75.0 \text{ mL} \end{array}$$
4. Chlorine gas occupies a volume of 25.0 mL at 300 K. What volume will it occupy at 600 K?

$$\begin{array}{ccc} V_1 & T_1 & T_2 \\ 25.0 \text{ mL} & 300 \text{ K} & 600 \text{ K} \end{array}$$
5. A sample of neon gas at $50^\circ C$ and a volume of 2.50 Liters is cooled to $25^\circ C$. What is the new volume?

$$\begin{array}{ccc} T_1 & V_1 & T_2 \\ 323 \text{ K} & 2.50 \text{ L} & 298 \text{ K} \end{array}$$
6. Fluorine gas at 300 K occupies a volume of 500 mL. To what temperature should it be lowered to bring the volume to 300 mL?

$$\begin{array}{ccc} T_1 & V_1 & V_2 \\ 300 \text{ K} & 500 \text{ mL} & 300 \text{ mL} \end{array}$$
7. Helium occupies a volume of 3.80 Liters at $-45^\circ C$. What volume will it occupy at $45^\circ C$?

$$\begin{array}{ccc} V_1 & T_1 & T_2 \\ 3.80 \text{ L} & 228 \text{ K} & 318 \text{ K} \end{array}$$
8. A sample of argon gas is cooled and its volume went from 380 mL to 250 mL. If its final temperature was $-55^\circ C$, what was its original temperature?

$$\begin{array}{ccc} V_1 & V_2 & T_2 \\ 380 \text{ mL} & 250 \text{ mL} & 218 \text{ K} \end{array}$$

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