

Why aren't the elements listed in alphabetical order?

Why aren't the elements listed in alphabetical order?

As charts go, the periodic table is a bit odd. It's not square. Large portions of the table appear to be missing at the top. It's not organized alphabetically so elements can be found easily. But to a chemist, the periodic table is a very powerful tool. The periodic table is organized by properties, both chemical and physical. Those properties relate to the electronic structure of the atoms of each element. In today's activity, you are going to study how the ground state electron configurations and structure of atoms are related to the shape and organization of the periodic table.

A blank periodic table grid with the following structure:

- s-block:** The first two columns (Groups 1 and 2).
- p-block:** The last six columns (Groups 13 to 18).
- d-block:** The ten columns between Groups 2 and 13 (Groups 3 to 10).
- f-block:** Two rows of seven columns each, located below the main body of the table, corresponding to the lanthanide and actinide series.

The grid is labeled with 's-block', 'p-block', 'd-block', and 'f-block' in the appropriate regions. The first column is numbered 1 through 7, and the first row of the f-block is numbered 1 through 7.

1. Obtain a card with electron configurations for your group as assigned by your teacher.
  - a. Record your team number\_\_\_\_\_ and team name \_\_\_\_\_
  - b. For each element on your card, write the end of the configuration (the underlined portion).
2. What is similar about the last orbital notation appearing at the end of the configuration for each element in your set?

# Breaking The Code Periodic Table Answers

**Marshall D. Abrams**



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