

Experiment

OBJECTIVE :

To analyse the given samples of commercial antacids by determining the amount of hydrochloric acid they can neutralize.

REQUIREMENTS :

Burettes, pipettes, titration flasks, measuring flasks, beakers, weight box, fractional weights, sodium hydroxide, sodium carbonate, hydrochloric acid, phenolphthalein.

PROCEDURE :

- 1) Prepare 1 litre of approximately N/10 HCl solution by diluting 10 ml of the concentrated acid for one litre.
- 2) Similarly, make 1 litre of approximately N/10 NaOH solution by dissolving 4.0g of NaOH to prepare one litre of solution.
- 3) Prepare N/10 Na_2CO_3 solution by weighing exactly 1.325 g of anhydrous sodium carbonate and then dissolving it in water to prepare exactly 0.25 litres (250 ml) of solution.
- 4) Standardize the HCl solution by titrating it against the standard Na_2CO_3 solution using methyl orange as indicator.
- 5) Similarly, standardize NaOH solution by titrating it against standardized HCl solution using phenolphthalein as indicator.
- 6) Powder the various samples of antacid tablets and weigh 1.0 g of each.
- 7) Add a specific volume of standardised HCl to each of the weighed sample is taken in conical flasks. The acid should be in slight excess, so that it can neutralize all the alkaline component of the tablet.
- 8) Add 2 drops of phenolphthalein and warm the flask till most of powder dissolves. Filter off the insoluble material.
- 9) Titrate this solution against the standardised NaOH solution, till a permanent pinkish tinge is obtained. Repeat this experiment with different antacids.

Effective Commercial Antacids Experiment

Kathy Dodds Tyner



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Laboratory Manual for Principles of General Chemistry Jo Allan Beran, 2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures

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Laboratory Experiments for General, Organic & Biochemistry Frederick A. Bettelheim, Joseph Marvin Landesberg, 1997 *Inquiry-based Experiments in Chemistry* Valerie Ludwig Lechtanski, 2000 Inquiry Based Experiments in Chemistry is an alternative to those cookbook style lab manuals providing a more accurate and realistic experience of scientific investigation and thought for the high school chemistry or physical science student

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Laboratory Exercises for Preparatory Chemistry Kathy Dodds Tyner, 1994-06 Laboratory Exercises for Preparatory Chemistry is the perfect complement to a one semester preparatory chemistry laboratory course

Tyner's manual emphasizes the application of chemistry and the principles of science to everyday life. The labs are directly applicable to the real world and often contain supplemental assignments that illustrate an application. *Chemistry of Common Substances* Lee R. Summerlin, 1979. *Bulletin of the South Carolina Academy of Science* South Carolina Academy of Science, 1986. List of members 1924-34 in v 1 p 21-26. *Psychopharmacology Abstracts*, 1965. *Chemical Engineering Education*, 1989. **Laboratory Handbook for General Chemistry** Norman E. Griswold, Conrad L. Stanitski, H. A. Neidig, James N. Spencer, 2006. The LABORATORY HANDBOOK FOR GENERAL CHEMISTRY helps students perform their laboratory work more effectively, efficiently, and safely. It is not a compilation of experimental procedures but rather, throughout three editions, it remains a how-to guide containing specific information about the basic equipment techniques and operations that are necessary for successful laboratory experiments. The importance of laboratory safety is stressed. Video demonstrations of a number of common laboratory techniques are an important feature of this Third Edition. The Handbook can be used in conjunction with CER modular experiments to support locally written experiments or to complement the techniques sections of commercial lab manuals. Laboratory Manual for Fundamentals of Chemistry 3/E Jo A. Beran, 1988. Antacids in the Eighties Fred Halter, 1982. **Introductory Chemistry** Wendy Gloffke, Doris Kimbrough, 2002. This lab manual helps students develop data acquisition, organization, and analysis skills while teaching basic techniques. Students construct their own data tables, answer conceptual questions, and make predictions before performing experiments. They also have the opportunity to visualize and describe molecular-level activity and explain the results.

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