

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

## CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY

### Solutions

#### Section 15.1 What are solutions?

In your textbook, read about the characteristics of solutions.

Use each of the terms below just once to complete the passage.

immiscible	liquid	soluble	solution
insoluble	miscible	solvent	solvent

Air is air! (1) solution of oxygen gas dissolved in nitrogen gas. The oxygen in air is the (2) solute and nitrogen is the (3) solvent. Because oxygen gas dissolves in a solvent, oxygen gas is air! (4) soluble substances. A substance that does not dissolve is (5) insoluble. (6) liquid solutions are the most common type of solutions. If one liquid is soluble in another liquid, such as acetic acid in water, the two liquids are (7) miscible. However, if one liquid is insoluble in another, the liquids are (8) immiscible.

Read about hydration in aqueous solutions in your textbook.

The diagram shows the hydration of solid sodium chloride to form an aqueous solution. Use the diagram to answer the following questions.



8. Hydration is a solution in which the solvent is water. What is solute?

Solution is the process of surrounding solute particles with solvent particles to form a solution.

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85

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## CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY

#### Section 15.1 continued

10. As sodium chloride dissolves in water, what happens to the sodium and chloride ions?

The sodium and chloride ions are separated and surrounded by the water molecules.

11. Explain the attraction of the water molecules around the sodium ions and chloride ions.

Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atoms). Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.

12. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? How do you know?

The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solution process to occur.

13. List three ways that the rate of solution may be increased.

stirring or shaking the solution, breaking the solute into smaller pieces, and heating the solvent

In your textbook, read about heat of solution, solubility, and factors that affect solubility.

For each statement below, write true or false.

True \_\_\_\_\_ 14. The overall energy change that occurs when a solution forms is called the heat of solution.

False \_\_\_\_\_ 15. Solubility is a measure of the maximum amount of solute that dissolves in a given amount of solvent at a specified temperature and pressure.

False \_\_\_\_\_ 16. Solubility continues to keep as the solubility rate is less than the crystals solubility rate.

True \_\_\_\_\_ 17. In a saturated solution, solubility and crystallization are in equilibrium.

True \_\_\_\_\_ 18. Additional solute can be dissolved in an unsaturated solution.

True \_\_\_\_\_ 19. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

86 Chemistry: Matter and Change • Chapter 15

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# Chemistry Study Guide For Content Mastery Answers

**Karin Nielsen-Saines**



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### **Table of Contents Chemistry Study Guide For Content Mastery Answers**

1. Understanding the eBook Chemistry Study Guide For Content Mastery Answers
  - The Rise of Digital Reading Chemistry Study Guide For Content Mastery Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Study Guide For Content Mastery Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Study Guide For Content Mastery Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Study Guide For Content Mastery Answers

- Personalized Recommendations
  - Chemistry Study Guide For Content Mastery Answers User Reviews and Ratings
  - Chemistry Study Guide For Content Mastery Answers and Bestseller Lists
5. Accessing Chemistry Study Guide For Content Mastery Answers Free and Paid eBooks
    - Chemistry Study Guide For Content Mastery Answers Public Domain eBooks
    - Chemistry Study Guide For Content Mastery Answers eBook Subscription Services
    - Chemistry Study Guide For Content Mastery Answers Budget-Friendly Options
  6. Navigating Chemistry Study Guide For Content Mastery Answers eBook Formats
    - ePub, PDF, MOBI, and More
    - Chemistry Study Guide For Content Mastery Answers Compatibility with Devices
    - Chemistry Study Guide For Content Mastery Answers Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Chemistry Study Guide For Content Mastery Answers
    - Highlighting and Note-Taking Chemistry Study Guide For Content Mastery Answers
    - Interactive Elements Chemistry Study Guide For Content Mastery Answers
  8. Staying Engaged with Chemistry Study Guide For Content Mastery Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Chemistry Study Guide For Content Mastery Answers
  9. Balancing eBooks and Physical Books Chemistry Study Guide For Content Mastery Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Chemistry Study Guide For Content Mastery Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Chemistry Study Guide For Content Mastery Answers
    - Setting Reading Goals Chemistry Study Guide For Content Mastery Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Chemistry Study Guide For Content Mastery Answers

- Fact-Checking eBook Content of Chemistry Study Guide For Content Mastery Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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