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- 2-1** (a) Stoichiometry is the description of the quantitative relationships among elements in a compound and among substances as they undergo chemical change.  
 (b) Composition stoichiometry describes the quantitative relationships among elements in compounds, e.g., in water,  $H_2O$ , there are 2 hydrogen atoms for every 1 atom of oxygen. Reaction stoichiometry describes the quantitative relationships among substances as they undergo chemical changes. (Reaction stoichiometry will be discussed in Chapter 3.)

The common ions for each formula unit is listed below:

(a)  $MgCl_2$  contains  $Mg^{2+}$  and  $Cl^-$  ions

(b)  $(NH_4)_2CO_3$  contains  $NH_4^+$  and  $CO_3^{2-}$  ions

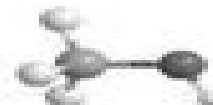
(c)  $Zn(NO_3)_2$  contains  $Zn^{2+}$  and  $NO_3^-$  ions

**2-5** Ethanol- $CH_3CH_2OH$



(space-filling; ball-and-stick)

Methanol- $CH_3OH$



(space-filling; ball-and-stick)

Both are composed of hydrogen, carbon, and oxygen. Both have an oxygen and hydrogen on the end. The ethanol molecule has an additional carbon and two hydrogens.

- 2-7** Organic compounds are those that contain carbon-to-carbon bonds, carbon-to-hydrogen bonds, or both. Organic formulas given in Table 2-1 include: acetic acid- $CH_3COOH$ , methane- $CH_4$ , ethane- $C_2H_6$ , propane- $C_3H_8$ , butane- $C_4H_{10}$ , pentane- $C_5H_{12}$ , benzene- $C_6H_6$ , methanol- $CH_3OH$ , ethanol- $CH_3CH_2OH$ , acetone- $CH_3COCH_3$ , diethyl ether- $CH_3CH_2COCH_2CH_3$ .

- 2-9** Compounds from Table 2-1 that contain only carbon and hydrogen and are not shown in Figure 1-5:

Compound	Ball and stick model	Compound	Ball and stick model
acetic acid- $CH_3COOH$		acetone- $CH_3COCH_3$	

# Chemistry Whitten Solution Manual

**Wendy L. Keeney-Kennicutt**



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### **Table of Contents Chemistry Whitten Solution Manual**

1. Understanding the eBook Chemistry Whitten Solution Manual
  - The Rise of Digital Reading Chemistry Whitten Solution Manual
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Whitten Solution Manual
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Whitten Solution Manual
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Whitten Solution Manual
  - Personalized Recommendations
  - Chemistry Whitten Solution Manual User Reviews and Ratings
  - Chemistry Whitten Solution Manual and Bestseller Lists

5. Accessing Chemistry Whitten Solution Manual Free and Paid eBooks
  - Chemistry Whitten Solution Manual Public Domain eBooks
  - Chemistry Whitten Solution Manual eBook Subscription Services
  - Chemistry Whitten Solution Manual Budget-Friendly Options
6. Navigating Chemistry Whitten Solution Manual eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Whitten Solution Manual Compatibility with Devices
  - Chemistry Whitten Solution Manual Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Whitten Solution Manual
  - Highlighting and Note-Taking Chemistry Whitten Solution Manual
  - Interactive Elements Chemistry Whitten Solution Manual
8. Staying Engaged with Chemistry Whitten Solution Manual
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Whitten Solution Manual
9. Balancing eBooks and Physical Books Chemistry Whitten Solution Manual
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Whitten Solution Manual
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Whitten Solution Manual
  - Setting Reading Goals Chemistry Whitten Solution Manual
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Whitten Solution Manual
  - Fact-Checking eBook Content of Chemistry Whitten Solution Manual
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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