

Thermochemistry Worksheet Key

1. How much heat, in calories, is released or absorbed if a 15.0 g sample of copper is converted from
(a) the solid to liquid state.

$$q = m \times \Delta H_{fus} = 15.0 \text{ g} \times 206 \frac{\text{J}}{\text{g}} \times \frac{1 \text{ cal}}{4.184 \text{ J}} = 738.5 = 739 \text{ cal}$$

So 739 cal is absorbed

- (b) the gas to the liquid state.

$$\begin{aligned} q &= m \times \Delta H_{con} = 15.0 \text{ g} \times -4.7 \frac{\text{kJ}}{\text{g}} \times \frac{1000 \text{ J}}{1 \text{ kJ}} \times \frac{1 \text{ cal}}{4.184 \text{ J}} = -16,849 \\ &= -1.68 \times 10^4 \text{ cal or } -16.8 \text{ kcal} \end{aligned}$$

So 16.8 kcal is released.

For copper m.p. = 1084 °C ; b.p. = 2927 °C; ΔH_{fus} = 206 J/g, ΔH_{vap} = 4.7 kJ/g

2. How much heat is absorbed or released if a 12.5 g sample of aluminum changes temperature from 25.0 to 87.5 °C? The specific heat capacity of aluminum is 0.900 J/g K

$$T_{final} = 87.5 + 273.15 = 360.65 \text{ K} \quad T_{initial} = 25.0 + 273.15 = 298.15 \text{ K}$$

$$\begin{aligned} q &= m \times c \times \Delta T = 12.5 \text{ g} \times 0.900 \frac{\text{J}}{\text{g} \cdot \text{K}} \times (360.65 - 298.15) \text{ K} \\ &= 12.5 \text{ g} \times 0.900 \frac{\text{J}}{\text{g} \cdot \text{K}} \times 62.5 \text{ K} = 703.1 = 703 \text{ J} \end{aligned}$$

3. What was the mass of a sample of water if the addition of 4.28×10^3 joules raised the temperature of the sample from 22.0 °C to 34.5 °C?

$$T_{final} = 34.5 + 273.15 = 307.65 \text{ K} \quad T_{initial} = 22.0 + 273.15 = 295.15 \text{ K}$$

T increases so endothermic so $q > 0$

$$q = m \times c \times \Delta T$$

$$\text{so } m = \frac{q}{c \times \Delta T} = \frac{4.28 \times 10^3 \text{ J}}{4.184 \frac{\text{J}}{\text{g} \cdot \text{K}} \times (307.65 - 295.15) \text{ K}} = \frac{4.28 \times 10^3 \text{ J}}{4.184 \frac{\text{J}}{\text{g} \cdot \text{K}} \times 12.5 \text{ K}} = 81.83 = 81.8 \text{ g}$$

Chemistry Thermochemistry Homework Packet Answers

Melissa May Batchelor



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General Chemistry Workbook Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material

Learning about Materials Ted Lister, Colin Osborne, 1998 This book presents useful material in the form of teacher's notes and photocopiable worksheets It will be useful at pre and post 16 level Synthesis and Characterization of Nitric Oxide-releasing Agents/polymers for Biomedical Applications Melissa May Batchelor, 2004

Publications of the National Bureau of Standards, 1979 Catalog United States. National Bureau of Standards, 1980

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Table of Contents Chemistry Thermochemistry Homework Packet Answers

1. Understanding the eBook Chemistry Thermochemistry Homework Packet Answers
 - The Rise of Digital Reading Chemistry Thermochemistry Homework Packet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Thermochemistry Homework Packet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Thermochemistry Homework Packet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Thermochemistry Homework Packet Answers

- Personalized Recommendations
 - Chemistry Thermochemistry Homework Packet Answers User Reviews and Ratings
 - Chemistry Thermochemistry Homework Packet Answers and Bestseller Lists
5. Accessing Chemistry Thermochemistry Homework Packet Answers Free and Paid eBooks
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 - Highlighting and Note-Taking Chemistry Thermochemistry Homework Packet Answers
 - Interactive Elements Chemistry Thermochemistry Homework Packet Answers
 8. Staying Engaged with Chemistry Thermochemistry Homework Packet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Thermochemistry Homework Packet Answers
 9. Balancing eBooks and Physical Books Chemistry Thermochemistry Homework Packet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Thermochemistry Homework Packet Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemistry Thermochemistry Homework Packet Answers
 - Setting Reading Goals Chemistry Thermochemistry Homework Packet Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemistry Thermochemistry Homework Packet Answers

- Fact-Checking eBook Content of Chemistry Thermochemistry Homework Packet Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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