

## CP Chemistry Unit 1 Test Review KEY

### Problems and questions:

- Round each of the following numbers to 3 significant figures:  
a) .09882 **.0988**      b) 123,459 **123,000**      c) 102.45 **102**      d)  $2.301 \times 10^{-2}$   **$2.30 \times 10^{-2}$**
- Calculate, use sig figs:  $(12.22 \times 14.02 \times 120.1 \times .555) / 10.00 =$  **1140**
- A sample of metal having a mass of 12.122 g is added to a graduated cylinder containing 35.2 mL of water. The water level rises to 38.4mL. Calculate the density, use correct significant figures.  
$$d = \frac{12.122}{3.2} = 3.8 \text{ g/mL}$$
- Calculate the % error for an experiment where the student measured 3.45 g and the actual value was 3.00 g?  
$$\% \text{ error} = \frac{(3.45 \text{ g} - 3.00 \text{ g})}{(3.00 \text{ g})} \times 100 = 15.0\%$$
- Write the symbol for the following ions: Mg losing 2 e<sup>-</sup>, Al losing 3 e<sup>-</sup>, N gaining 3 e<sup>-</sup>, S gaining 2 e<sup>-</sup>  
**Mg<sup>2+</sup>, Al<sup>3+</sup>, N<sup>3-</sup>, S<sup>2-</sup>**
- a) 2 b) 3 c) 12 d) 5
- List three points of Dalton's atomic theory.  
A. Elements made of tiny particles, called atoms.  
B. Atoms of a given element are identical; the atoms of different elements are different.  
C. In chemical reactions, atoms are combined, separated, or rearranged. Atoms cannot be subdivided, created, or destroyed.  
D. Atoms combine with atoms, in small whole numbers, to form compounds.
- Describe Thompson's experiment with the cathode ray tube, what did he discover?  
Cathode ray tube consists of a cathode (negative) and anode (positive) side. The "ray" could be bent with magnetic or electric fields. The "ray" bent toward a positive EM field and away from a negative EM field. Discovered electron and charge to mass ratio of electron.
- Describe Millikan's Oil Drop experiment. What exactly did he determine?  
He suspended a micro oil droplet in an electric field and was able to determine the actual charge of an electron to be  $-1.6 \times 10^{-19} \text{ C}$ . Also, determined the mass of an electron.
- Explain Rutherford's Gold Foil experiment. What were three observations and three conclusions?  
He fired alpha particles at gold foil and observed: 1) most went straight through, 2) small number showed a slight deflection, 3) very few bounced straight back. He concluded: 1) mass of atom is in the nucleus, 2) nucleus is positively charged, 3) atom mostly empty space.
- What is an isotope? **Atoms of the same element that differ by # of electrons.**
- Write the symbol of the most abundant copper isotope if the average atomic mass is 63.55?  **$^{64}_{29}\text{Cu}$**
- Write the nuclear symbol for an isotope with Z=40, and 45 neutrons?  **$^{85}_{40}\text{Zr}$**
- Consider the diagram and match the letters with the appropriate particle  
**A= electron, B= neutron, C= proton**
- Place the forms of hydrogen (H, H<sup>+</sup>, H<sup>-</sup>, H<sub>2</sub>) in the table.



Atom	Cation	Anion	Molecule
<b>H</b>	<b>H<sup>+</sup></b>	<b>H<sup>-</sup></b>	<b>H<sub>2</sub></b>

# Chemistry Unit 1 Solutions Test Review Answers

**Liying Dong**



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SILBERBERG,2003      **Study guide to accompany Drew H. Wolfe: General, organic and biological chemistry** Daniel C. Pantaleo,1986      *Journal of Applied Chemistry* ,1872      *El-Hi Textbooks in Print* ,1984      *Introduction to General, Organic, and Biochemistry* Morris Hein,1997 The authors recognize that both science and mathematics may be daunting subjects for many students taking this course With this in mind they have anticipated where students might stumble and have paced and organized this text to help them through Their goal is to make the material interesting and relevant so students understand the basic chemical principles related to their career The authors emphasize problem solving and provide a range of practice exercises As in previous editions the text first presents the basic concepts of general chemistry and then moves into organic and biochemistry In this edition the first two sections have been revised primarily to improve explanations and include new pedagogical features The biochemistry portion has been thoroughly updated to include coverage of many recent developments and emerging technologies in the field      *Lithographic Technology in Transition* Ervin A. Dennis,Olusegun Odesina,Daniel G. Wilson,1996 Designed for the serious graphic communications student this text examines both the traditional and cutting edge technologies of lithography Technology business and career topics are discussed in depth as the text assists the reader in making the transition from traditional lithography to the high technology printing processes of the present and the future      **Introductory Chemistry** Wendy Gloffke,Doris Kimbrough,2002 This lab manual helps students develop data acquisition organization and analysis skills while teaching basic techniques Students construct their own data tables answer conceptual questions and make predictions before performing experiments They also have the opportunity to visualize and describe molecular level activity and explain the results      **Foundations Science** Janice Harriet Weiss,Contemporary,1993-04 Contemporary s Foundation series provides thorough coverage of basic skills at reading levels 4 6 Gives students meaningful contexts for learning Makes materials easy to understand Provides students with the opportunity to create essay answers and practice the steps of the writing process Post tests assess skills proficiency upon completing each books Evaluation Charts target and prescribe areas for needed practice Thorough coverage of the writing process analyzing the essay writing the essay and mechanics      **Fundamentals of Chemistry** David E. Goldberg,2001      **Medical Review** ,1907      **Basic Medical Lab Techniques-Iml 4e** Estridge,Ruth Reynolds,2000-05      *Nuclear Science Abstracts* ,1972

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