

Chemistry – Unit 5 Worksheet 1

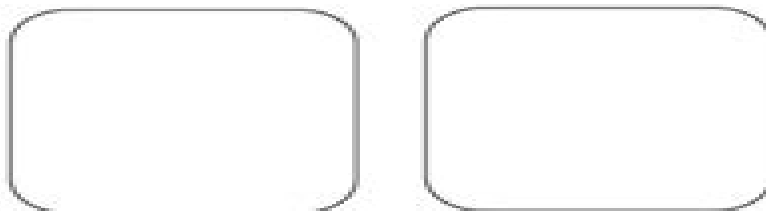
Molar Masses of the Elements

Relative Mass From Gases

We have established that the combining ratio of gases can be explained if two assumptions are made:

1. Equal volumes of gases contain the same number of molecules at the same pressure and temperature.
2. Some pure elemental gases are clustered into pairs to form diatomic molecules.

Use these assumptions and particle diagrams to explain the fact that the density of oxygen gas at standard temperature and pressure is 1.43 g/liter, whereas the density of hydrogen gas under these conditions is 0.089 g/liter. How many times more massive is one molecule of oxygen than one molecule of hydrogen?



You shouldn't conclude that chemists were able to determine the molar masses of all the elements using this technique. Measurements of the density of the gaseous phase of many of the elements are difficult, if not impossible. However, we are going to see that chemists could use another tool – the *percent composition of compounds* to determine molar masses.

Relative Mass From Compounds

Many substances combine with oxygen to form a type of compound called an oxide. In the previous unit we saw that such combinations often occur in multiple proportions. John Dalton made the assumption that the lowest ratio was a 1:1 combination of elements. For now, we will make a similar assumption. *We may have to re-examine this assumption later.*

In Unit 4 you examined % composition data of several compounds. For example, 100 g of the simplest oxide of carbon contains 42.9g of C and 57.1 g of O. To be able to compare masses of elements to the mass of oxygen, it is important to use the same mass of oxygen in all of the ratios. Using a proportion, one can show that 75.0 g of C combines with 100 g of oxygen.

$$\frac{42.9\text{gC}}{57.1\text{gO}} = \frac{75.0\text{gC}}{100\text{gO}}$$

Chemistry Unit 5 1 Answers

Michal Rosen-Zvi

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located on the right side of the slide.

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Ebook: Chemistry: The Molecular Nature of Matter and Change Silberberg, 2015-01-16 Ebook Chemistry The Molecular Nature of Matter and Change **Exercises for the Anatomy & Physiology Laboratory** Erin C. Amerman, 2019-02-01 This concise inexpensive black and white manual is appropriate for one or two semester anatomy and physiology laboratory courses It offers a flexible alternative to the larger more expensive laboratory manuals on the market This streamlined manual shares the same innovative activities based approach as its more comprehensive full color counterpart Exploring Anatomy Physiology in the Laboratory 3e Basic Concepts of Chemistry Leo J. Malone, Theodore O. Dolter, 2011-12-27 The 9th edition of Malone s Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter WileyPLUS sold separately from text **Solving Problems in Chemistry** Rod O'Connor, Charles Mickey, Alton Hassell, 1977 **Aim@AIIMS-JEE Mains** R.CHANDRAN.S.V., 2021-07-13 NEET JEE mains NCERT Based *Chemistry for the Allied Health Sciences* Roger D. Bauer, Robert L. Loesch, 1980 Think Level 4 Teacher's Book Brian Hart, Herbert Puchta, Jeff Stranks, Peter Lewis-Jones, 2016-09-08 Challenge and inspire your teenage learners to think beyond language Think is a fresh vibrant and upbeat course designed to engage teenage learners and make them think The Teacher s Book provides full lesson support for this level including teacher s notes audio scripts answer keys suggestions for extra activities and lesson options for mixed ability classes plus tips on how to get the best out of Think s array of digital tools and components Clinical Chemistry Theresa Dean, Sheryl Whitlock, 1997 This complete procedural manual offers laboratory technicians hands on coverage of all aspects of clinical chemistry Using numerous real world case studies it fully examines test principles and demonstrates correct test techniques and procedures Effective quality control and safety procedures are incorporated throughout 55 illus International Critical Tables of Numerical Data, Physics, Chemistry and Technology: X-ray data National Research Council (U.S.), 1929 **International Critical Tables of Numerical Data, Physics, Chemistry and Technology** National Research Council (U.S.), 1929 State Board Examination Questions and Answers of Twenty-three States , 1907 **International Critical Tables of Numerical Data, Physics, Chemistry and Technology** , 1929 Cehmistry Textbook for College and University USA Ibrahim Sikder, 2023-06-04 Cehmistry Textbook USA **The Publishers' Trade List Annual** , 1904 Michigan School Moderator , 1901 *Christian Advocate* , 1908

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