

Name Tara ShettyDate 9/24/15 Pd 5**Chemistry – Unit 4 Worksheet 3**

Use the following information about the masses of elements in each pair of compounds to help you suggest formulas that account for these ratios.

1. Compounds of carbon and oxygen

Compound A: 57.1 g O / 42.9 g C

Compound B: 72.7 g O and 27.3 C

$$A = \frac{57.1}{42.9} = 1 \frac{1}{3}$$

$$B = \frac{72.7}{27.3} = 2 \frac{2}{3}$$

a. Determine the value of the ratio $\frac{\text{mass O}}{\text{mass C}}$ in each compound. A $1 \frac{1}{3}$ B $2 \frac{2}{3}$

b. How does the mass ratio for compound B compare to that in compound A?

compound B is larger than compound A

c. Express these ratios as improper fractions. $\frac{57.1}{42.9} + \frac{72.7}{27.3} \left(\frac{4}{3} + \frac{8}{3} \right)$

d. For each hypothesis, sketch particle diagrams for the compounds of A and B that account for these mass ratios. Write the formula for the compound in each diagram.

Hypothesis 1

Atoms of C and O have the same mass

A $\frac{4}{3}$  C_3O_4 B $\frac{8}{3}$  C_8O_3 **Hypothesis 2**

Atoms of O are heavier than C atoms by the ratio in compound A.

A

CO

B

 CO_2

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