

Chemistry Notes for class 12 Chapter 2 Solutions

Solution is a homogeneous mixture of two or more substances in same or different physical phases. The substances forming the solution are called components of the solution. On the basis of number of components a solution of two components is called binary solution.

Solute and Solvent

In a binary solution, solvent is the component which is present in large quantity while the other component is known as solute.

Classification of Solutions

(A) Following types of solutions are seen on the basis of physical state of solute and solvent.

S.No.	Solute	Solvent	Examples
Solid solutions			
1	Solid	Solid	Alloys
2	Liquid	Solid	Hydrated salts. Amalgam of Hg with Na
3	Gas	Solid	Dissolved gases in mineral
Liquid solutions			
4	Solid	Liquid	Salt/sugar solution in water
5	Liquid	Liquid	Alcohol in water
6	Gas	Liquid	Aerated drinks, O_2 in water
Gaseous solutions			
7	Solid	Gas	Iodine vapour in air
8	Liquid	Gas	Water vapour in air
9	Gas	Gas	Air ($O_2 + N_2$)

[if water is used as a solvent, the solution is called aqueous solution and if not, the solution is called non-aqueous solution.]

(B) Depending upon the amount of solute dissolved in a solvent we have the following types of solutions:

(i) **Unsaturated solution** A solution in which more solute can be dissolved without raising temperature is called an unsaturated solution.

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