

Chemistry C2 1.1-2.5

Atomic Structure

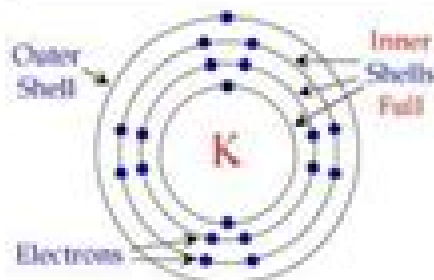
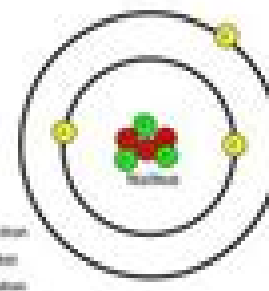
The element **Lithium** has a symbol of **Li**.

3 electrons (atomic number) and an atomic mass of 7

3
Li
6.94

Atoms are made of **protons**, **neutrons** and **electrons**.

Protons have a **positive +** charge and electrons have a **negative -** charge. Atoms are arranged in the periodic table in order of their atomic number. Neutrons have **no** electric charge so are **neutral**.



The electrons in an atom are arranged in **shells**.

Atoms with the same number of electrons in their outer shell belong in the same group of the periodic table. The number of electrons in the outer shell of an atom determines the way that the atom behaves in chemical reactions.

Chemical Bonding

Elements react to form compounds by **gaining** or **losing** electrons or by **sharing** electrons.

The elements in Group 1 react with the elements in Group 7 because group 1 needs to **lose** an electron from its outer shell and group 7 need to **gain** an electron to have a full outer shell.

This can be shown by using **dot and cross diagrams**.

Ionic Bonding

Ionic compounds are held together by strong forces between the oppositely charged ions. This is called **ionic bonding**.

Group 2 & Group 6 elements can form ionic compounds. This happens between a metal and non-metal. These form giant structures of ions in a lattice.

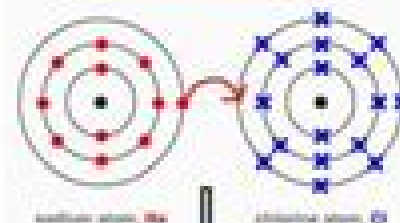
Covalent Bonding

This takes place between **non-metals**. Covalent bonds are formed when atoms **share** electrons. Many substances contain covalent bonds consist of molecules, but some have giant covalent structures. The electrons are shared so there is an outer shell of 8 electrons.

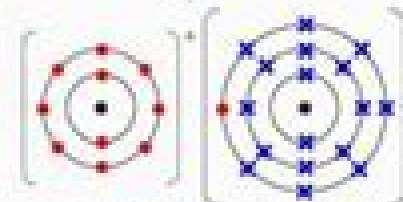
Remember: An outer shell which does not have a full shell of 8 electrons is **unstable**. So all elements are looking to become **stable**.



ionic bonding in sodium chloride



sodium atom, Na chlorine atom, Cl



sodium ion, Na⁺ chloride ion, Cl⁻

free electrons from outer shells of metal atoms



Metallic Bonding

The atoms (or ions) in metals are arranged in regular layers. The positive ions in metals are held together by electrons from the outer shell of each metal atom. These form **giant** structures. The diagram on the left shows this.

Ionic compounds - it takes a lot of **energy** to break the bonds which hold a giant ionic lattice together. Very **high melting points** and are solid at room temperatures. They conduct electricity when melt or dissolve them in water because their ions can then move freely.

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