

Colligative (collective) properties are vapor pressure lowering, boiling point elevation (ΔT_b), freezing point depression (ΔT_f), and osmotic pressure (Π). **Colligative properties depend only on the number of solute particles and not their chemical identity.** This is best, specifically, the properties of a pure solvent compare with those of a solution:

- The vapor pressure of the solution is lower.
- The boiling point of the solution is higher.
- The freezing (or melting) point of the solution is lower.
- The solution causes osmosis, the flow of solvent from a solution of lower solute concentration to one of higher solute concentration a semipermeable membrane.

Below are the key equations you need for solving problems related to colligative properties.

Key Equations for Colligative Properties

Raoult's Law: Relationship between the Vapor Pressure of a Solution (P_{solution}), the Mole Fraction of the Solvent (χ_{solvent}), and the Vapor Pressure of the Pure Solvent ($P^{\circ}_{\text{solvent}}$)

$$P_{\text{solution}} = \chi_{\text{solvent}} P^{\circ}_{\text{solvent}}$$

$$\chi_{\text{solvent}} = \frac{P_{\text{solution}}}{P^{\circ}_{\text{solvent}} + P_{\text{solute}}}$$

The Vapor Pressure of a Solution Containing Two Volatile Components

$$P_A = \chi_A P^{\circ}_A \quad P_B = \chi_B P^{\circ}_B$$

$$P_{\text{total}} = P_A + P_B$$

Relationship between Osmotic Pressure (Π), Molarity (M), the Ideal Gas Constant (R), and Temperature (T , in K)

$$\Pi = MRT$$

$$M = \frac{n_{\text{solute}}}{V_{\text{solution}}} \quad L = \text{liters} / \text{mol} \cdot K$$

The dependence of ΔT_f and ΔT_b on the Molarity (m) and Freezing Point Depression Constant (K_f) and Boiling Point Elevation Constant (K_b) respectively:

$$\frac{\Delta T_f}{\Delta T_b} = \frac{m \cdot K_f}{m \cdot K_b} = m = \frac{\text{mol (solute)}}{\text{kg (solvent)}}$$

Boiling-Point and Freezing-Point Constants

Solvent	Freezing point (°C)	K_f (K·kg·mol ⁻¹)	Boiling point (°C)	K_b (K·kg·mol ⁻¹)
acetone	-93.85	2.40	56.3	1.71
benzene	5.5	5.12	80.1	2.53
carbon tetrachloride	-23	29.8	76.5	4.93
cyclohexane	6.5	20.1	80.7	2.79
naphthalene	80.3	6.94	217.7	5.80
phenol	43	7.27	182	3.84
water	0	1.86	100.0	0.51

An important trait in colligative properties is, the larger the number of the nodes, the stronger the effect. For example, dissolving 2 moles of methanol (nonelectrolyte) will not affect the colligative properties of the solution as much as 2 moles of NaCl (electrolyte) would. This is because NaCl dissociates into ions, so 2 moles of NaCl will produce 4 moles of particles (ions).



Electrolytes have stronger effect on colligative properties of solutions.

Therefore, we divide our summary in two parts: colligative properties of nonelectrolytes solutions and those of electrolyte solutions.

Nonvolatile Nonelectrolyte Solutions

Vapor Pressure Lowering Example: Calculate the vapor pressure of a solution at 25°C that is made by adding 47.5 g of glucose ($C_6H_{12}O_6$) to 340 g of water. The vapor pressure of pure water at 25°C is 23.8 torr.

The vapor pressure of the solution is calculated by the following formula:

$$P_{\text{solution}} = \chi_{\text{solvent}} P^{\circ}_{\text{solvent}}$$

We have the vapor pressure of pure water: $P^{\circ}(\text{H}_2\text{O}) = 23.8 \text{ Torr}$, so the only missing part is the mole fraction of the solvent (water in this case) which is the ratio of the moles of solvent over the total moles of the solvent and the solute.

Calculate the moles and plug the numbers on the equation for the mole fraction of H_2O :

$$\chi_{\text{H}_2\text{O}} = \frac{n(\text{H}_2\text{O})}{n(\text{H}_2\text{O}) + n(\text{C}_6\text{H}_{12}\text{O}_6)} = \frac{18.9 \text{ mol}}{(18.9 + 0.298) \text{ mol}} = 0.986$$

The vapor pressure of the solution would be: $P_{\text{solution}} = 0.986 \times 23.8 \text{ torr} = 23.5 \text{ torr}$

Freezing Point Depression Example: Using the appropriate data in the table, determine the freezing point depression of the solution that contains 24.1 g urea (NH_2CONH_2) in 485 mL of water.

The freezing point depression is calculated by the following formula: $\Delta T_f = m \cdot K_f$

So, we need to first determine the moles of urea, then calculate the molarity and use it in the equation for the ΔT_f .

$$n(\text{NH}_2\text{CONH}_2) = 24.1 \text{ g} \times \frac{1 \text{ mol}}{60.1 \text{ g}} = 0.401 \text{ mol}$$

$$m(\text{H}_2\text{O}) = 485 \text{ mL} \times \frac{1.00 \text{ g}}{1 \text{ mL}} \times \frac{1 \text{ kg}}{1000 \text{ g}} = 0.485 \text{ kg}$$

$$m = \frac{0.401 \text{ mol}}{0.485 \text{ kg}} = 0.827 \text{ m}$$

$$\Delta T_f = K_{fm} = 1.86^\circ\text{C} \cdot \text{m} \times 0.827 \text{ m} = 0.591^\circ\text{C}$$

Chemistry Study Guide Calculations Involving Colligative Properties

Dixie Goss



Chemistry Study Guide Calculations Involving Colligative Properties:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03

The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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