PROBLEMS pg. 69 #7, 11, 13, 15, 17, 20, 21, 22, 27, 28, 30, 31, 33, 35, 38, 40, 41, 43, 45, 47, 49, 55, 56, 57, 58, 60, 61, 64, 67, 69, 71, 87

7 The diameter of a neutral belium atom = 1 × 10<sup>2</sup> pm. Suppose that we could line up belium atoms side by side in contact with one another. Approximately how many atoms would it take to make the distance from end to end 1 cm<sup>2</sup>

First, convert 1 cm to picometers.  $1 \text{ cm} \times \frac{0.01 \text{ m}}{1 \text{ cm}} \times \frac{1 \text{ pm}}{1 \times 10^{-12} \text{ m}} = 1 \times 10^{10} \text{ pm}$ 

? He atoms =  $(1 \times 10^{10} \text{ pm}) \times \frac{1 \text{ He atom}}{1 \times 10^{2} \text{ pm}} = 1 \times 10^{8} \text{ He atoms}$ 

- What do we call atoms of the same elements with different mass numbers? Isotopes.
- What is the mass number of an iron atom that has 28 neutrons?
  For iron, the atomic number Z is 26. Therefore the mass number A is:
  A = 26 + 28 = 54
- 15 For the following atoms, determine the number of protons and neutrons.

Isotope	He	Hic	28 Mg	25 Mig	22.13	no Bar	ne Pe
No. Protons	.2	2	1.2	12	2.2	3.5	78
No. Neutrons	1	2	12	1.3	26	44	117

17 Give the symbol for the following isotopes:

(a) Z = 11 A = 23  ${}_{11}^{13}$ Na (b) Z = 28 A = 64  ${}_{28}^{64}$ Ni

20 State two differences between a metal and a nonmetal.

Metals conduct heat and electricity, are shiny & malleable; nonmetals are poor conductors, generally dull and brittle.

21 Write the names and symbols for four elements in each of the following categories:

(a) nonemetal: phosphorous P, bromine Br, neon, Ne, carbon C

(b) metal: iron Fe, cesium Cs, iridium, Ir, praseodymium Pr

(c) metalloid: silicon Si, germanium Ge, boron B, antimony Sb

22 Define, with two examples, the following terms:

(a) Alkali metals: (1" column) Sodium, rubidium;

(b) alkaline earth metals: (2nd column) barium, beryllium;

(c) halogens: (7th column) fluorine, iodine;

(d) noble gases: (8th column) xenon, krypton

27 What is the difference between an atom and a molecule?

An atom is the smallest representative unit of an element.

A molecule is the smallest representative unit of a (non-ionic) compound, composed of two or more atoms chemically bonded.

28 What are allotropes? Allotropes are different physical forms of the same element. Examples? graphite and diamond are two allotropes of carbon.

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