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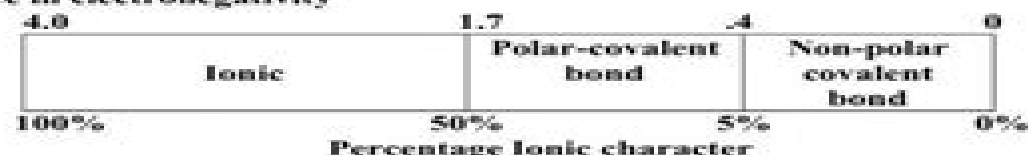
Electronegativity: A property of an atom which increases with its tendency to attract the electrons of a bond.

Examples: The chlorine atom has a higher electronegativity than the hydrogen atom, so the bonding electrons will be closer to the Cl than to the H in the HCl molecule.

		H 2.1														
1A	2A											3A	4A	5A	6A	7A
Li 1.0	Be 1.5	Electronegativity										B 2.0	C 2.5	N 3.0	O 3.5	F 4.0
Na 1.0	Mg 1.2	3B	4B	5B	6B	7B	8B					Al 1.5	Si 1.8	P 2.1	S 2.5	Cl 3.0
K 0.9	Ca 1.0	Sc 1.3	Ti 1.4	V 1.5	Cr 1.6	Mn 1.6	Fe 1.7	Co 1.7	Ni 1.8	Cu 1.8	Zn 1.6	Ga 1.7	Ge 1.9	As 2.1	Se 2.4	Br 2.8
Rb 0.9	Sr 1.0	Y 1.2	Zr 1.3	Nb 1.6	Mo 1.6	Tc 1.7	Ru 1.8	Rh 1.8	Pd 1.8	Ag 1.6	Cd 1.6	In 1.6	Sn 1.8	Sb 1.9	Te 2.1	I 2.5
Cs 0.8	Ba 1.0	La 1.1	Hf 1.3	Ta 1.4	W 1.5	Re 1.7	Os 1.9	Ir 1.9	Pt 1.8	Au 1.9	Hg 1.7	Tl 1.6	Pb 1.7	Bi 1.8	Po 1.9	At 2.1

<1.0	1.5–1.9	2.5–2.9	4.0
1.0–1.4	2.0–2.4	3.0–3.9	

Difference in electronegativity



Bonding between	More electronegative element and value	Less electronegative element and value	Difference in electronegativity	Bond Type
Sulfur and Hydrogen				
Sulfur and cesium				
Chlorine and bromine				
Calcium and chlorine				
Oxygen and hydrogen				
Nitrogen and hydrogen				
Iodine and iodine				
Copper and sulfur				
Hydrogen and fluorine				
Carbon and oxygen				

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Worksheet 24

Yi-Tong Ma



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