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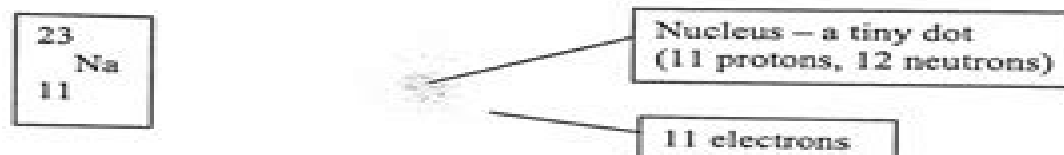
Subatomic Particles

Particle	Symbol	Relative Charge	Absolute Mass	Relative Mass
electron	e^{-}	-1	$9.109 \times 10^{-31} \text{ kg}$	0
proton	p^{+}	+1	$1.673 \times 10^{-27} \text{ kg}$	1
neutron	n^0	0	$1.675 \times 10^{-27} \text{ kg}$	1

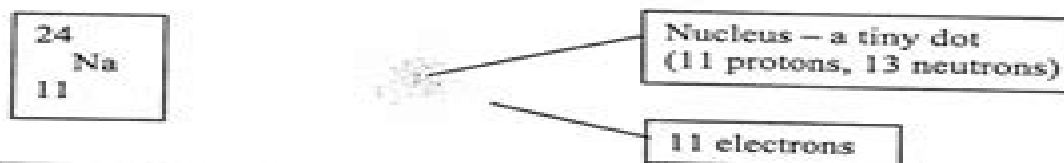
Model: Two Isotopes of Sodium

The diagrams below show representations of sodium isotopes. [Note: the diameter of an atom is about 10,000 times larger than the diameter of the atomic nucleus so the relative sizes of the atom and the nucleus are not accurately depicted in these diagrams.]

Isotope 1



Isotope 2

**Key Questions**

1. What information is provided by the atomic number, Z?

of protons

2. What information is provided by the mass number, A?

protons + neutrons

Chemistry Unit 2 Packet Key

Mary Millhollon, Richard H. Langley

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the authors' names.

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