

4.7 - REVIEW – Unit 4 – Chem. Compds & Bonding

- Atoms bond to become **STABLE** which is to say it will gain a stable and full **OCTET**.
- Ionic bonding occurs between **METALS** and **NONMETALS**. They are held together due to **OPPOSITE** charges attracting.
- In general, **METALS** are more likely to be oxidized because they have **FEWER** valence electrons while **NONMETALS** are more likely to be reduced because they have **MORE** valence electrons.
- When calcium reacts with bromine, the calcium becomes **OXIDIZED** (oxidized, reduced) and the bromine becomes **REDUCED** (oxidized, reduced). The oxidizing agent is **BROMINE**, and the reducing agent is **CALCIUM**.
- A metallic bond is the attraction between the **NUCLEUS** and the **NEGATIVE** sea of **ELECTRONS**. Which occurs between two **METALS**.
- When a transition metal forms a bond, it loses electrons from **THE S-SHELL**. For example, Zn, with electron configuration **[Ar] 4s² 3d¹⁰** will form an ion with a charge of **+2**. The ion's configuration will be **[Ar] 4s⁰ 3d¹⁰**.
- Write the equation for the formation of the following ions (elements becoming stable).



Chemistry Unit 4 Test Answers

Sasha Hu

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