

Chemistry – Unit 5

Worked Example 5 – Gases Again

1. A sealed system (fixed amount) containing gas was investigated. The pressure of the gas at 10°C was 100 kPa. The temperature is 10°C. The table below is the total number of molecules of the gas at different temperatures. What is the pressure at the next level?

Temperature / °C	Number of molecules
10	100
20	120
30	140
40	160
50	180

2. A 100 mL volume of carbon monoxide gas is present at 10°C and 100 kPa. At what volume would the pressure of this gas be 100 kPa? Assume temperature does not change.

3. A small bubble containing methane is taken from the temperature is 10°C and the pressure is 100 kPa. In the bubble's surface, where the temperature is 10°C and the pressure is 100 kPa, calculate the mass of methane gas in the bubble. Is the bubble getting bigger or smaller?

4. Three gases are mixed in a 1.00 L container. The partial pressures of CO_2 is 100 kPa, O_2 is 100 kPa and 100 kPa, and there is 100 kPa. What is the pressure of the mixture of gases?

5. What is the percentage by mass of the gases that are mixed?

Chemistry Unit 9 Ws 2 Answers

M Carnoy



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