

Aspects	Intramolecular forces	Intermolecular forces
Nature of forces	Strong	Weak
Types of Interaction	Atomic bonding	Molecular bonding
Determination	Chemical properties	Physical properties
Relative strength	Generally stronger	Generally weaker
Range of action	Within a single molecule	Between different molecules
Influence on reactivity	Effect chemical reactions	Effects molecular interactions
Importance	For understanding molecular structure and chemical reactivity	For explaining physical properties and intermolecular interactions
Examples	<ul style="list-style-type: none"> • Ionic bonds • Covalent bonds • Coordinate colvanet bonds • Metallic bonds 	<ul style="list-style-type: none"> • Dipole-dipole interactions • Ion-dipole interactions • Instantaneous dipole-Ion dipole interactions • Van der Waals forces • Hydrogen bonding

Chemists Distinguish Between Intermolecular And Intramolecular Forces

Robert K. Wismer



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Edexcel AS Chemistry Student Unit Guide: Rod Beavon, 2012-03-09 Student Unit Guides are perfect for revision Each guide is written by an examiner and explains the unit requirements summarises the relevant unit content and includes a series of specimen questions and answers There are three sections to each guide Introduction includes advice on how to use the guide an explanation of the skills being tested by the assessment objectives an outline of the unit or module and depending on the unit suggestions for how to revise effectively and prepare for the examination questions Content Guidance provides an examiner's overview of the module's key terms and concepts and identifies opportunities to exhibit the skills required by the unit It is designed to help students to structure their revision and make them aware of the concepts they need to understand the exam and how they might analyse and evaluate topics Question and Answers sample questions and with graded answers which have been carefully written to reflect the style of the unit All responses are accompanied by commentaries which highlight their respective strengths and weaknesses giving students an insight into the mind of the examiner

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An Introduction to Chemistry Michael Mosher, Paul Kelter, 2023-03-18 This textbook is written to thoroughly cover the topic of introductory chemistry in detail with specific references to examples of topics in common or everyday life It provides a major overview of topics typically found in first year chemistry courses in the USA The textbook is written in a conversational question based format with a well defined problem solving strategy and presented in a way to encourage readers to think like a chemist and to think outside of the box Numerous examples are presented in every chapter to aid students and provide helpful self learning tools The topics are arranged throughout the textbook in a traditional approach to the subject with the primary audience being undergraduate students and advanced high school students of chemistry

Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001

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