

Classifying Chemical Reactions

Analyzing and Predicting Products



Introduction

The power of chemical reactions to transform our lives is visible all around us—in our homes, in our cars, even in our bodies. Chemists try to make sense of the great variety of chemical reactions the same way that biologists organize their knowledge of life, by sorting reactions into groups and classifying them. Classifying chemical reactions allows us to predict what chemical reactions will occur when different substances are mixed.

Concepts

- Chemical reactions
- Single vs. double replacement
- Combination vs. decomposition
- Combustion reactions

Background

A chemical reaction is defined as any process in which one or more substances are converted into new substances with different properties. Chemical reactions change the identity of the reacting substance(s) and produce new substances. Observing the properties of the reactants and products is therefore a key step in identifying chemical reactions. Some of the observations that may be associated with a chemical reaction include: (1) release of a gas; (2) formation of a precipitate; (3) color changes; (4) temperature changes; (5) emission or absorption of light. As these observations suggest, chemical reactions can be dynamic and exciting events. The essence of any chemical reaction—reactants being transformed into products—is summarized in the form of a chemical equation. Consider the reaction represented by Equation 1, the burning of natural gas (methane, CH_4) in a laboratory burner.



The reactants—on, more specifically, their formulas—are written on the left side of the equation, the products on the right side of the equation. An arrow represents the direction of the reaction and is read as “yields” or “produces.” Other symbols are used to describe the physical state of the reactants and products and to describe the reaction conditions (see Table 1 on page 2).

Chemical reactions may be classified by considering the number of reactants and products in the reactions, the physical or chemical nature of the reactants and products, and the rearrangement of atoms in the conversion of the reactants into products (see Table 2 on page 2).

Different textbooks may classify reactions differently. Some authors treat all single replacement reactions within a much larger group of oxidation-reduction reactions. Other authors refer to double-replacement reactions as metathesis or double displacement reactions. Finally, a few authors use neither designation and classify double replacement reactions as either precipitation or acid-base reactions.

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ML Yell



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Dissertation Abstracts International ,1985-07 *Ceramic Abstracts* American Ceramic Society,1993 *Chemical Reaction Networks* Oleg N. Temkin,Andrew V. Zeigarnik,D. G. Bonchev,2019-12-02 Over the last decade increased attention to reaction dynamics combined with the intensive application of computers in chemical studies mathematical modeling of chemical processes and mechanistic studies has brought graph theory to the forefront of research It offers an advanced and powerful formalism for the description of chemical reactions and their intrinsic reaction mechanisms *Chemical Reaction Networks A Graph Theoretical Approach* elegantly reviews and expands upon graph theory as applied to mechanistic theory chemical kinetics and catalysis The authors explore various graph theoretical approaches to canonical representation numbering and coding of elementary steps and chemical reaction mechanisms the analysis of their topological structure the complexity estimation and classification of reaction mechanisms They discuss topologically distinctive features of multiroute catalytic and noncatalytic and chain reactions involving metal complexes With it s careful balance of clear language and mathematical rigor the presentation of the authors significant original work and emphasis on practical applications and examples *Chemical Reaction Networks A Graph Theoretical Approach* is both an outstanding reference and valuable tool for chemical research

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