

Addition Compounds

Double Salt.

Complex Salt.

* Carnallite :-



* Mohr's Salt :-



* Potash Alum :-



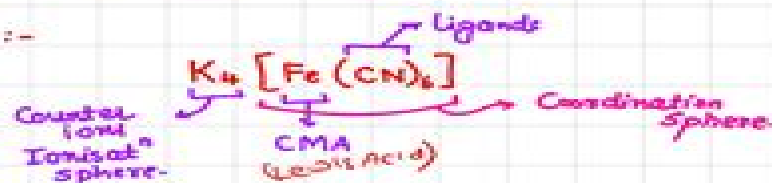
* d-block elements.

↳ * Small Size,

* high ϕ Nature.

* Vacant Orbitals

Note :-



* **Denticity** :- No. of e^- pairs which can be donated by a ligand to metal ion.

* In any ligand if total no. of atoms are 3 or less than 3 \rightarrow then denticity is always "1".

Stable \rightarrow free

CLASSIFICATION OF LIGANDS :-

A) Anionic

SO_4^{2-} = Sulfate

CO_3^{2-} = Carbonate

Cl^- = Chloride

OH^- = Hydroxide

F^- = Fluoride

NH_2^- = Amide

NH^{2-} = Imide

N_3^- = Azide

N^{3-} = Nitride

O^{2-} = Oxide

O_2^- = Superoxide

O_2^{2-} = Peroxide

CN^- = Cyanide

NC^- = Isocyanide

NO_2^- = Nitrite - N

ONO^- = Nitrite - O

OCN^- = Cyanate - O

NCO^- = Cyanate - N

SCN^- = Thiocyanate - S

NCS^- = Thiocyanate - N

$\text{C}_2\text{O}_4^{2-}$ = Oxalate (2)

Gly = Glycinate (2)

dim = dimethylglycinate (2)

glyox = glyoximate (2)

EDTA^{4-} = Ethylene diamine tetraacetate. (6)

B) Neutral :-

H_2O :- Aqua

NH_3 :- Amine

CO :- Carbonyl

NO :- Nitrosyl

PH_3 :- Phosphine

PPh_3 :- Triphenyl Phosphine

Py :- Pyridine.

en :- ethylenediamine.

pn :- propylene diamine

tn :- trimethylene diamine

dien :- diethylene triamine

taien :- triethylene tetraamine.

C) Cationic :-

NO^+ = Nitrosonium

$\text{NH}_2^+ - \text{NH}_3^+$:- Hydrazonium.

Jugaad :-

$\rightarrow \text{Pt/Pd}$

always d^8

$\text{CN} = 4$

O.S. = +2

$\text{CN} = 6$

O.S. = +4.

\rightarrow When Fe & NO both +1.

$\text{NO} = +1$

$\rightarrow \text{Sc Ti V Cr Mn Fe Co Ni Cu Zn} \}^{+2}$

3d = 1 2 3 4 5 6 7 8 9 10

Similar for below 4d & 5d !

$\rightarrow d^1/d^2/d^3 \xrightarrow[6 \text{ ligands}]{\text{SL/OL}} d^2sp^3$
 $\text{UPE}^- = 1 \ 2 \ 3$ $\text{CN} = 6$ always Ioc LSC Paramag.

$\rightarrow d^2/d^3/d^4 \xrightarrow[6 \text{ ligands}]{\text{SL/OL}} sp^3d^2$
 $\text{UPE}^- = 2 \ 1 \ 0$ $\text{CN} = 6$ always ooc HSC Para Para dia.

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