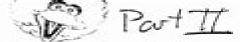
CLASSIFYING CHEMICAL REACTIONS

H,SO,

14. SO₃ + H₂O →



Classify the following reactions as synthesis, decomposition, single replacement or double replacement. AACCLERS

1. $2KClO_3 \rightarrow 2KCl + 3O_2$ Decomposition 2. $HCl + NaOH \rightarrow NaCl + H_2O$ Double Replacement 3. $Mg + 2HCl \rightarrow MgCl_2 + H_2$ Single Replacement 4. $2H_2 + O_2 \rightarrow 2H_2O$ Synthesis 5. $2Al + 3NlBr_2 \rightarrow 2AlBr_3 + 3Nl$ Single Replacement 6. $4Al + 3O_2 \rightarrow 2Al_2O_3$ Synthesis 7. $2NaCl \rightarrow 2Na + Cl_2$ Decomposition 8. $CaCl_2 + F_2 \rightarrow CaF_2 + Cl_2$ Single Replacement 9. $AgNO_3 + KCl \rightarrow AgCl + KNO_3$ Double Replacement 10. $N_2 + 3H_2 \rightarrow 2NH_3$ Synthesis 11. $2H_2O_2 \rightarrow 2H_2O + O_2$ Decomposition 12. $(NH_2)_2SO_4 + Bac(NO_3)_2 \rightarrow BasO_4 + 2NH_4NO_3$ Double Replacement 13. $Mgl_2 + Br_2 \rightarrow MgBr_2 + l_2$ Single Replacement	repidcemeni.	ANSW	ERS/
3. Mg + 2HCI → MgCl2 + H2 Single Replacement. 4. 2H2 + O2 → 2H2O Synthesis 5. 2AI + 3NIBr2 → 2AIBr3 + 3NI Single Replacement 6. 4AI + 3O2 → 2AI2O3 Synthesis 7. 2NaCI → 2Na + Cl2 Decomposition 8. CaCl2 + F2 → CaF2 + Cl2 Single Replacement 9. AgNO3 + KCI → AgCI + KNO3 Double Replacement 10. N2 + 3H2 → 2NH3 Synthesis 11. 2H2O2 → 2H2O + O2 Decomposition 12. (NH2)2SO4 + Ba(NO3)2 → BaSO4 + 2NH4NO3 Double Replacement	1. 2KCIO ₃ → 2KCI + 3O ₂	De con	~position
4. $2H_2 + O_2 \rightarrow 2H_2O$ Synthesis 5. $2AI + 3NIBI_2 \rightarrow 2AIBI_3 + 3NI$ Single Replacement 6. $4AI + 3O_2 \rightarrow 2AI_2O_3$ Synthesis 7. $2NaCI \rightarrow 2Na + CI_2$ By composition 8. $CaCI_2 + F_2 \rightarrow CaF_2 + CI_2$ 9. $AgNO_3 + KCI \rightarrow AgCI + KNO_3$ Double Replacement 10. $N_2 + 3H_2 \rightarrow 2NH_3$ Synthesis 11. $2H_2O_2 \rightarrow 2H_2O + O_2$ Decomposition 12. $(NH_2)_2SO_4 + Ba(NO_3)_2 \rightarrow BaSO_4 + 2NH_3NO_3$ Double Replacement	2. HCl + NaOH → NaCl + H₂O	Double	Replacement
5. 2AI + 3NIBr ₂ → 2AIBr ₃ + 3NI Single Replacement 6. 4AI + 3O ₂ → 2AI ₂ O ₃ Synthesis 7. 2NOCI → 2NO + CI ₂ Decomposition 8. COCI ₂ + F ₂ → COF ₂ + CI ₂ Single Replacement 9. AGNO ₃ + KCI → AGCI + KNO ₃ Double Replacement 10. N ₂ + 3H ₂ → 2NH ₃ Synthesis 11. 2H ₂ O ₂ → 2H ₂ O + O ₂ Decomposition 12. (NH ₂ , SO ₄ + BO(NO ₃) ₂ → BOSO ₄ + 2NH ₄ NO ₃ Double Replacement	3. Mg + 2HCl → MgCl ₂ + H ₂	Single	Replacement.
6. 4AI + 30, - 2AI2O, 7. 2NacI - 2Na + CI2 8. CacI2 + F2 - CaF2 + CI2 9. AgNO3 + KCI - AgCI + KNO3 10. N2 + 3H2 - 2NH3 11. 2H2O2 - 2H2O + O2 12. (NH2)2SO4 + Ba(NO3)2 - BaSO4 + 2NH4NO3 Double Replacement Decomposition 12. (NH2)2SO4 + Ba(NO3)2 - BaSO4 + 2NH4NO3 Double Replacement Decomposition	4. 2H ₂ + O ₂ → 2H ₂ O	Synth	esis
7. 2NaCl - 2Na + Cl2 Be composition 8. CaCl2 + F2 - CaF2 + Cl2 9. AgNO3 + KCl - AgCl + KNO3 10. N2 + 3H2 - 2NH3 11. 2H2O2 - 2H2O + O2 De composition De composition De composition De composition De composition	5. 2AI + 3NIBr₂ → 2AIBr₃ + 3NI	Single	Replacement
8. Cacl ₂ + F ₂ → CaF ₂ + Cl ₂ Single Replacement 9. AgNO ₃ + KCl → AgCl + KNO ₃ Double Replacement 10. N ₂ + 3H ₂ → 2NH ₃ Synthesis 11. 2H ₂ O ₂ → 2H ₂ O + O ₂ Decomposition 12. (NH ₂) ₂ SO ₄ + Ba(NO ₃) ₂ → BaSO ₄ + 2NH ₄ NO ₃ Double Replacement	6. 4AI + 3O₂ → 2AI₂O₃	Synth	e515
9. AgNO3 + KCI - AgCI + KNO3 Double Replacement 10. N2 + 3H2 - 2NH3 Synthesis 11. 2H2O2 - 2H2O + O2 Decomposition 12. (NH2)2SO4 + Ba(NO3)2 - BaSO4 + 2NH2NO3 Double Replacement	7. 2NaCl → 2Na + Cl ₂	Decom	position
10. N ₂ + 3H ₂ → 2NH ₃ Synthesis 11. 2H ₂ O ₂ → 2H ₂ O + O ₂ Decomposition 12. (NH ₂) ₂ SO ₄ + Ba(NO ₃) ₂ → BaSO ₄ + 2NH ₄ NO ₅ Double Replacement	8. $CaCl_2 + F_2 \rightarrow CaF_2 + Cl_2$	Single	Replacement
11. 2H2O2 - 2H2O + O2 De composition 12. (NH2)2SO4 + BO(NO3)2 - BOSO4 + 2NH4NO, Double Replacement	9. AgNO ₃ + KCI → AgCI + KNO ₃	Double	Replacement
12. (NH.)250, + Ba(NO), - Baso, + 2NH, NO, Double Replacement	10. N ₂ + 3H ₂ → 2NH ₃	Synth	resis
	11. 2H ₂ O ₂ → 2H ₂ O + O ₂	Decomp	position
13. Mgl2 + Br2 - MgBr2 + 12 Single Replacement	12. (NH ₄) ₂ SO ₄ + Ba(NO ₃) ₂ → BaSO ₄ +	+ 2NH4NO, DO	uble Replacement
	13. Mgl ₂ + Br ₂ → MgBr ₂ + l ₂	Single	Replacement

Synthesis

Lifeliqe

Real World Examples of Four Types of Chemical Reactions, James W Smith developed Real World Examples of Four Types of Chemical Reactions a physical science and chemistry lesson for 9th grade students. The students use the Internet to locate information about industrial pharmaceutical and environmental operations that use examples of the four general classes of chemical reactions. These reactions are synthesis decomposition single replacement reactions and double replacement reactions. The Louisiana Challenge Grant provides the lesson online as part of the Louisiana Challenge collection of activities for the K 12 classroom.

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