

## Analyzing and Predicting Products



## Concepts

- ## Background

$$4\text{HCl(g)} + \text{NO}_2\text{(g)} \rightarrow \text{Cl}_2\text{(g)} + 2\text{H}_2\text{O(g)}$$

**Formalizing  $\mathcal{L}$**

Chemical reactions may be classified by considering the number of reactants and products in the reaction, the physical or chemical nature of the reactants and products, and the rearrangement of atoms in the conversion of the reactants into products (see Table 2 on page 28).

Different textbooks may classify reactions differently. Some authors treat all single replacement reactions within a much larger group of oxidation-reduction reactions. Other authors refer to double-replacement reactions as *synthesis* or *double displacement* reactions. Finally, a few authors use neither designation and classify double replacement reactions as *other precipitation* or *acid-base* reactions.

# Classifying Chemical Reactions Lab Flinn Answers

**Ingo Waldemar Dagobert Hackh**



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**Chemical Reactions and Their Equations** Ingo Waldemar Dagobert Hackh, 1921

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