

# Stoichiometry Introduction

Stoichiometry is the math side of Chemistry! It relates different quantities of products and reactants in chemical reactions. Stoichiometry relies on balanced equations to find the mole ratios used in most calculations. A mole ratio is the ratio of moles (coefficients) of one reactant/product involved in a reaction to another. We can use mole ratios to convert from moles of one substance to moles of another. But we can also convert between grams and moles.

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## Example 1: Finding mole ratio

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There are 5 moles of  $\text{O}_2$  and 2 moles of  $\text{CO}_2$ . This means that the mole ratios of  $\text{O}_2$  to  $\text{CO}_2$  are  $\frac{5 \text{ mol } \text{O}_2}{2 \text{ mol } \text{CO}_2}$  and  $\frac{2 \text{ mol } \text{CO}_2}{5 \text{ mol } \text{O}_2}$ .

1. Find all the mole ratios for the equation  $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ .

2. Find all the mole ratios for the equation  $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$ .

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## Example 2: mole to mole conversion

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If we start with 3 moles of  $\text{C}_2\text{H}_6$ , how many moles of  $\text{H}_2\text{O}$  will be produced?

$$3 \text{ mol } \text{C}_2\text{H}_6 \times \frac{4 \text{ mol } \text{H}_2\text{O}}{2 \text{ mol } \text{C}_2\text{H}_6}$$

The mol  $\text{C}_2\text{H}_6$  cancel, leaving you with the mol of  $\text{H}_2\text{O}$ .

$$3 \text{ mol } \text{C}_2\text{H}_6 \times \frac{4 \text{ mol } \text{H}_2\text{O}}{2 \text{ mol } \text{C}_2\text{H}_6} = 12 \text{ mol } \text{H}_2\text{O}$$

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Use the following equations for #3-#5:  $\text{Ba}^{2+} + 2\text{F}^- \rightarrow \text{BaF}_2$

3. How many moles of  $\text{KF}$  are produced when 5 mol of  $\text{Ba}^{2+}$  react?

# Chapter Stoichiometry Packet

**James N. Spencer, George M.  
Bodner, Lyman H. Rickard**



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