

## Chapter 9 Stoichiometry Chapter Review Answers

- Mole Ratios
  - Mole ratios are conversion factors that relate the number of moles of one chemical to another chemical involved in a chemical reaction.
  - They are formed from the coefficients found in a balanced chemical equation.
- Mole Ratios as conversion factors
  - $\frac{2 \text{ mol Ca}}{1 \text{ mol O}_2} \cdot \frac{1 \text{ mol O}_2}{2 \text{ mol Ca}} \cdot \frac{2 \text{ mol CaO}}{2 \text{ mol Ca}} = \frac{2 \text{ mol CaO}}{2 \text{ mol CaO}} = \frac{1 \text{ mol O}_2}{1 \text{ mol CaO}} \cdot \frac{2 \text{ mol CaO}}{1 \text{ mol O}_2}$ . Remember the ratios can be reduced as well!!!
  - $\frac{1 \text{ mol Mg}}{2 \text{ mol HF}} \cdot \frac{2 \text{ mol HF}}{1 \text{ mol Mg}} \cdot \frac{1 \text{ mol MgF}_2}{1 \text{ mol Mg}} = \frac{1 \text{ mol MgF}_2}{1 \text{ mol Mg}} \cdot \frac{1 \text{ mol Mg}}{1 \text{ mol Mg}} \cdot \frac{1 \text{ mol Mg}}{1 \text{ mol Mg}} = \frac{1 \text{ mol MgF}_2}{1 \text{ mol Mg}} \cdot \frac{1 \text{ mol Mg}}{1 \text{ mol Mg}} \cdot \frac{1 \text{ mol Mg}}{1 \text{ mol Mg}} = \frac{1 \text{ mol MgF}_2}{1 \text{ mol Mg}}$
- $\frac{165.99 \text{ g Na}_2\text{CO}_3}{1 \text{ mol Na}_2\text{CO}_3} \cdot \frac{1 \text{ mol Na}_2\text{CO}_3}{165.99 \text{ g Na}_2\text{CO}_3} \cdot \frac{74.10 \text{ g Ca(OH)}_2}{1 \text{ mol Ca(OH)}_2} \cdot \frac{1 \text{ mol Ca(OH)}_2}{74.10 \text{ g Ca(OH)}_2} = \frac{40.00 \text{ g NaOH}}{1 \text{ mol NaOH}} \cdot \frac{1 \text{ mol NaOH}}{40.00 \text{ g NaOH}} \cdot \frac{108.99 \text{ g CaCO}_3}{1 \text{ mol CaCO}_3} \cdot \frac{1 \text{ mol CaCO}_3}{108.99 \text{ g CaCO}_3}$
- Molar Mass
  - Definition: the mass of one mole of a substance
  - Molar mass is used as a conversion factor in stoichiometric calculations to convert mass of a substance to moles of the substance or convert moles of a substance to mass of the substance.
- Hydrogen + Oxygen  $\rightarrow$  Water
  - 5.0 mol H<sub>2</sub>
  - 2.5 mol O<sub>2</sub>
- Ethane combustion reaction
  - 15.8 mol O<sub>2</sub>
  - 9.00 mol CO<sub>2</sub> & 13.5 mol H<sub>2</sub>O
- 575 g Na & 886 g Cl<sub>2</sub>
- Iron production reaction
  - 75.1 mol CO
  - 50.1 mol Fe & 75.1 mol CO<sub>2</sub>
- 87,390 g CO or  $8.739 \times 10^5 \text{ g CO}$  & 12,610 g H<sub>2</sub> or  $1.261 \times 10^4 \text{ g H}_2$
- Lightning reaction of N<sub>2</sub> & O<sub>2</sub>
  - 1,100 g NO<sub>2</sub> or  $1.10 \times 10^3 \text{ g NO}_2$
  720. g NO or  $7.20 \times 10^2 \text{ g NO}$
- CO<sub>2</sub> removal in spacecraft
  - 42.04 mol NaOH
  - 21.02 mol Na<sub>2</sub>CO<sub>3</sub> & 21.02 mol H<sub>2</sub>O
- Double replacement AgNO<sub>3</sub> & NaBr
  - 463 g NaBr
  - 845 g AgBr
- Soda-acid fire extinguisher
  - 3.058 mol NaHCO<sub>3</sub>
  - 3.058 mol CO<sub>2</sub>, 1.529 mol Na<sub>2</sub>SO<sub>4</sub>, 3.058 mol H<sub>2</sub>O
- Sulfuric acid-sodium hydroxide neutralization reaction
  - H<sub>2</sub>SO<sub>4</sub> + 2 NaOH  $\rightarrow$  Na<sub>2</sub>SO<sub>4</sub> + 2 H<sub>2</sub>O
  - 37 g H<sub>2</sub>SO<sub>4</sub>
  - 53 g Na<sub>2</sub>SO<sub>4</sub> & 14 g H<sub>2</sub>O
- Copper & silver nitrate reaction
  - 0.0104 mol Cu(NO<sub>3</sub>)<sub>2</sub>
  - 0.0104 mol Cu & 0.0209 mol AgNO<sub>3</sub>
- Aspirin production
  - 13.5 kg C<sub>6</sub>H<sub>5</sub>O<sub>4</sub>
  - 7.66 kg C<sub>6</sub>H<sub>5</sub>O<sub>3</sub>
  - 4.29 L HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>
- Ideal stoichiometry calculations do not account for factors that can affect the relative amounts of product chemicals formed in reactions. Ideal stoichiometric calculations are based on the amounts of reactants and products under ideal conditions. Real stoichiometric calculations account for actual conditions under which the reaction occurs.
- The limiting reactant is the chemical in the reaction in least abundance for the reaction; it will be the chemical reactant that runs out first in the reaction. The limiting reactant **limits** the amount of chemical product(s) that can be formed. The excess reactant(s) will not be used up completely in the reaction; they will be leftover.

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**Kaplan Test Prep**



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**EBOOK: GENERAL CHEMISTRY, THE ESSENTIAL CONCEPTS** CHANG,2013-01-07 EBOOK GENERAL CHEMISTRY THE ESSENTIAL CONCEPTS **Complete Preparation for the MCAT** Williams & Wilkins Review,1998-04 Here is the most respected test prep book for the Medical College Admission Test you can buy featuring an active learning approach for a better understanding of the exam s content and a better chance for success Unique to this guide are coverage of all recent changes in the MCAT plus a step by step plan for sharpening cognitive skills developing problem solving skills and critical thinking This thorough guide replaces expensive test preparation courses while giving students exactly what they need to get ready for the MCAT Holt Chemistry Ralph Thomas Myers,2004 **DAT 2017-2018 Strategies, Practice & Review with 2 Practice Tests** Kaplan Test Prep,2016-10-04 Kaplan s DAT 2017 2018 Strategies Practice Review features the realistic practice test taking strategies and expert guidance you need to score higher on the Dental Admissions Test Our comprehensive subject review and test blueprint will help you face Test Day with confidence The Best Review Two full length online practice tests More than 600 practice questions for every subject with detailed answers and explanations 12 page tear out full color study sheets for quick review on the go A guide to the current DAT Blueprint so you know exactly what to expect on Test Day Comprehensive review of all of the content covered on the DAT Biology General Chemistry Organic Chemistry Perceptual Ability Reading Comprehension Quantitative Reasoning Kaplan s proven strategies for Test Day success Expert Guidance Kaplan s expert psychometricians ensure our practice questions and study materials are true to the test We invented test prep Kaplan www kaptest com has been helping students for almost 80 years Our proven strategies have helped legions of students achieve their dreams **OAT 2017-2018 Strategies, Practice & Review with 2 Practice Tests** Kaplan Test Prep,2016-10-04 Kaplan s OAT 2017 2018 Strategies Practice Review provides the content review test taking strategies and realistic practice you need to get the OAT results you want Updated for the latest test changes OAT 2017 2018 is your guide to facing Test Day with confidence The Best Review Two full length online practice tests More than 600 practice questions for every subject with detailed answers and explanations 16 page tear out full color study sheets for quick review on the go A guide to the current OAT Blueprint so you know exactly what to expect on Test Day Comprehensive review of all of the content covered on the OAT Biology General Chemistry Organic Chemistry Reading Comprehension Physics Quantitative Reasoning Kaplan s proven strategies for Test Day success Expert Guidance Kaplan s expert psychometricians ensure our practice questions and study materials are true to the test We invented test prep Kaplan www kaptest com has been helping students for almost 80 years Our proven strategies have helped legions of students achieve their dreams Chemical Principles Peter Atkins,Loretta Jones,2007-08 Written for calculus inclusive general chemistry courses Chemical Principles helps students develop chemical insight by showing the connections between fundamental chemical ideas and their applications Unlike other texts it begins with a detailed picture of the atom then builds toward

chemistry's frontier continually demonstrating how to solve problems, think about nature and matter, and visualize chemical concepts as working chemists do. Flexibility in level is crucial and is largely established through clearly labeling, separating, and boxing the calculus coverage in the text. Instructors have the option of whether to incorporate calculus in the coverage of topics. The multimedia integration of Chemical Principles is more deeply established than any other text for this course. Through the unique eBook, the comprehensive Chemistry Portal, Living Graph icons that connect the text to the Web, and a complete set of animations, students can take full advantage of the wealth of resources available to them to help them learn and gain a deeper understanding.

**Kaplan PCAT 2016-2017 Strategies, Practice, and Review with 2 Practice Tests**  
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