

# Stoichiometry

## Section 11.1 Defining Stoichiometry

pages 368–372

### Practice Problems

pages 371–372

1. Interpret the following balanced chemical equations in terms of particles, moles, and mass. Show that the law of conservation of mass is observed.



1 molecule  $\text{N}_2$  + 3 molecules

$\text{H}_2 \rightarrow$  2 molecules  $\text{NH}_3$

1 mole  $\text{N}_2$  + 3 moles  $\text{H}_2 \rightarrow$  2 moles  $\text{NH}_3$

Mass:

$\text{N}_2$ : (2 mol)(14.007 g/mol) = 28.014 g

$3\text{H}_2$ : (6 mol)(1.008 g/mol) = 6.048 g

$2\text{NH}_3$ : (2 mol)(14.007 g/mol) +  
(6 mol)(1.008 g/mol) = 34.062 g

28.014 g  $\text{N}_2$  + 6.048 g  $\text{H}_2 \rightarrow$  34.062 g  $\text{NH}_3$

34.062 g reactants = 34.062 g products



1 molecule  $\text{HCl}$  + 1 formula unit  $\text{KOH} \rightarrow$

1 formula unit  $\text{KCl}$  + 1 molecule  $\text{H}_2\text{O}$

1 mole  $\text{HCl}$  + 1 mole  $\text{KOH} \rightarrow$

1 mole  $\text{KCl}$  + 1 mole  $\text{H}_2\text{O}$

Mass:

$\text{HCl}$ : (1 mol)(1.008 g/mol) +  
(1 mol)(35.453 g/mol) = 36.461 g

$\text{KOH}$ : (1 mol)(39.098 g/mol) +  
(1 mol)(15.999 g/mol) + (1 mol)(1.008 g/mol)  
= 56.105 g

$\text{KCl}$ : (1 mol)(39.098 g/mol) +  
(1 mol)(35.453 g/mol) = 74.551 g

$\text{H}_2\text{O}$ : (2 mol)(1.008 g/mol) +  
(1 mol)(15.999 g/mol) = 18.015 g

36.461 g  $\text{HCl}$  + 56.105 g  $\text{KOH} \rightarrow$   
74.551 g  $\text{KCl}$  + 18.015 g  $\text{H}_2\text{O}$

92.566 g reactants = 92.566 g products



2 atoms  $\text{Mg}$  + 1 molecule

$\text{O}_2 \rightarrow$  2 formula units  $\text{MgO}$

2 moles  $\text{Mg}$  + 1 mole  $\text{O}_2 \rightarrow$  2 moles  $\text{MgO}$

Mass:

$2\text{Mg}$ : (2 mol)(24.305 g/mol) = 48.610 g

$\text{O}_2$ : (2 mol)(15.999 g/mol) = 31.998 g

$2\text{MgO}$ : (2 mol)(24.305 g/mol) +  
(2 mol)(15.999 g/mol) = 80.608 g

48.610 g  $\text{Mg}$  + 31.998 g  $\text{O}_2 \rightarrow$  80.608 g  $\text{MgO}$

80.608 g reactants = 80.608 g products

2. **Challenge** For each of the following, balance the chemical equation; interpret the equation in terms of particles, moles, and mass; and show that the law of conservation of mass is observed.



2 atoms  $\text{Na}$  + 2 molecules  $\text{H}_2\text{O} \rightarrow$

2 formula units  $\text{NaOH}$  + 1 molecule  $\text{H}_2$

2 mol  $\text{Na}$  + 2 mol  $\text{H}_2\text{O} \rightarrow$

2 mol  $\text{NaOH}$  + 1 mol  $\text{H}_2$

Mass:

$2\text{Na}$ : (2 mol)(22.990 g/mol) = 45.980 g

$2\text{H}_2\text{O}$ : (4 mol)(1.008 g/mol) +  
(2 mol)(15.999 g/mol) = 36.030 g

$2\text{NaOH}$ : (2 mol)(22.990 g/mol) +  
(2 mol)(15.999 g/mol) + (2 mol)(1.008 g/mol)  
= 79.994 g

$\text{H}_2$ : (2 mol)(1.008 g/mol) = 2.016 g

45.980 g  $\text{Na}$  + 36.030 g  $\text{H}_2\text{O} \rightarrow$

79.994 g  $\text{NaOH}$  + 2.016 g  $\text{H}_2$

82.01 g reactants = 82.01 g products

# Chapter Stoichiometry Section Answers

**Robinson**



## Chapter Stoichiometry Section Answers:

*Programmed Problem-solving for First-year Chemistry* Arnold B. Loebel, 1983      **Chemical Problem-solving by Dimensional Analysis** Arnold B. Loebel, 1978      Answer Guide for Chemistry, 2nd Edition Leonard W. Fine, 1978      *Basic Concepts of Chemistry* Leo J. Malone, Theodore Dolter, 2008-12-03 Engineers who need to have a better understanding of chemistry will benefit from this accessible book It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features Each section focuses on the development and assessment of one or two specific objectives Within each section a specific objective is included an anticipatory set to orient the reader content discussion from established authors and guided practice problems for relevant objectives These features are followed by a set of independent practice problems The expanded Making it Real feature showcases topics of current interest relating to the subject at hand such as chemical forensics and more medical related topics Numerous worked examples in the text now include Analysis and Synthesis sections which allow engineers to explore concepts in greater depth and discuss outside relevance      EBOOK: GENERAL CHEMISTRY, THE ESSENTIAL CONCEPTS CHANG, 2013-01-07 EBOOK GENERAL CHEMISTRY THE ESSENTIAL CONCEPTS      **OAT 2017-2018 Strategies, Practice & Review with 2 Practice Tests** Kaplan Test Prep, 2016-10-04 Kaplan s OAT 2017 2018 Strategies Practice Review provides the content review test taking strategies and realistic practice you need to get the OAT results you want Updated for the latest test changes OAT 2017 2018 is your guide to facing Test Day with confidence The Best Review Two full length online practice tests More than 600 practice questions for every subject with detailed answers and explanations 16 page tear out full color study sheets for quick review on the go A guide to the current OAT Blueprint so you know exactly what to expect on Test Day Comprehensive review of all of the content covered on the OAT Biology General Chemistry Organic Chemistry Reading Comprehension Physics Quantitative Reasoning Kaplan s proven strategies for Test Day success Expert Guidance Kaplan s expert psychometricians ensure our practice questions and study materials are true to the test We invented test prep Kaplan www kaptest com has been helping students for almost 80 years Our proven strategies have helped legions of students achieve their dreams      **Complete Preparation for the MCAT** Williams & Wilkins Review, 1998-04 Here is the most respected test prep book for the Medical College Admission Test you can buy featuring an active learning approach for a better understanding of the exam s content and a better chance for success Unique to this guide are coverage of all recent changes in the MCAT plus a step by step plan for sharpening cognitive skills developing problem solving skills and critical thinking This thorough guide replaces expensive test preparation courses while giving students exactly what they need to get ready for the MCAT      **Open-Ended Problems** James Patrick Abulencia, Louis Theodore, 2015-03-27 This is a unique book with nearly 1000 problems and 50 case studies on open ended problems in every key topic in chemical engineering that helps to better prepare chemical engineers for the future The term open ended problem basically describes an approach to the solution of a problem and or situation for which

there is not a unique solution The Introduction to the general subject of open ended problems is followed by 22 chapters each of which addresses a traditional chemical engineering or chemical engineering related topic Each of these chapters contain a brief overview of the subject matter of concern e g thermodynamics which is followed by sample open ended problems that have been solved by the authors employing one of the many possible approaches to the solutions This is then followed by approximately 40 45 open ended problems with no solutions although many of the authors solutions are available for those who adopt the book for classroom or training purposes A reference section is included with the chapter s contents Term projects comprised of 12 additional chapter topics complement the presentation This book provides academic industrial and research personnel with the material that covers the principles and applications of open ended chemical engineering problems in a thorough and clear manner Upon completion of the text the reader should have acquired not only a working knowledge of the principles of chemical engineering but also and more importantly experience in solving open ended problems What many educators have learned is that the applications and implications of open ended problems are not only changing professions but also are moving so fast that many have not yet grasped their tremendous impact The book drives home that the open ended approach will revolutionize the way chemical engineers will need to operate in the future

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## **Table of Contents Chapter Stoichiometry Section Answers**

1. Understanding the eBook Chapter Stoichiometry Section Answers
  - The Rise of Digital Reading Chapter Stoichiometry Section Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter Stoichiometry Section Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter Stoichiometry Section Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter Stoichiometry Section Answers
  - Personalized Recommendations
  - Chapter Stoichiometry Section Answers User Reviews and Ratings
  - Chapter Stoichiometry Section Answers and Bestseller Lists
5. Accessing Chapter Stoichiometry Section Answers Free and Paid eBooks
  - Chapter Stoichiometry Section Answers Public Domain eBooks
  - Chapter Stoichiometry Section Answers eBook Subscription Services
  - Chapter Stoichiometry Section Answers Budget-Friendly Options
6. Navigating Chapter Stoichiometry Section Answers eBook Formats

- ePub, PDF, MOBI, and More
- Chapter Stoichiometry Section Answers Compatibility with Devices
- Chapter Stoichiometry Section Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter Stoichiometry Section Answers
  - Highlighting and Note-Taking Chapter Stoichiometry Section Answers
  - Interactive Elements Chapter Stoichiometry Section Answers
- 8. Staying Engaged with Chapter Stoichiometry Section Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter Stoichiometry Section Answers
- 9. Balancing eBooks and Physical Books Chapter Stoichiometry Section Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter Stoichiometry Section Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter Stoichiometry Section Answers
  - Setting Reading Goals Chapter Stoichiometry Section Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter Stoichiometry Section Answers
  - Fact-Checking eBook Content of Chapter Stoichiometry Section Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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