

Name \_\_\_\_\_ Class \_\_\_\_\_ Date \_\_\_\_\_

## Assessment

**Chapter Test****Stoichiometry**

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- \_\_\_\_\_ 1. For the reaction  $\text{SO}_3 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{SO}_4$ , calculate the percentage yield if 500. g of sulfur trioxide react with excess water to produce 575 g of sulfuric acid.
- 82.7%
  - 88.3%
  - 91.2%
  - 93.9%
- \_\_\_\_\_ 2. In the reaction  $2\text{Al}_2\text{O}_3 \rightarrow 4\text{Al} + 3\text{O}_2$ , what is the mole ratio of aluminum to oxygen?
- 10:6
  - 3:4
  - 2:3
  - 4:3
- \_\_\_\_\_ 3. Fewer steps are required to solve stoichiometry problems when the reactant is given in
- grams and the product is sought in grams.
  - moles and the product is sought in moles.
  - grams and the product is sought in liters.
  - liters and the product is sought in number of atoms.
- \_\_\_\_\_ 4. Which of the following mathematical expressions correctly states the relationship among percentage yield, actual yield, and theoretical yield?
- $$\text{actual yield} = \frac{\text{percentage yield}}{\text{theoretical yield}} \times 100$$
  - $$\text{percentage yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$
  - $$\text{theoretical yield} = \frac{\text{actual yield}}{\text{percentage yield}} \times 100$$
- i
  - ii
  - iii
  - Both (b) and (c)

# Chapter Stoichiometry Test

**John T. Moore, Richard H. Langley**



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